

Exercise 1A

Name : _____

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1.2 Solving Quadratic Equations Using the Factor Method

Key Concepts and Formulae

1. An equation in the form $ax^2 + bx + c = 0$, where a , b and c are real numbers and $a \neq 0$, is a quadratic equation in one unknown x .
2. For any two real numbers a and b , if $ab = 0$, then $a = 0$ or $b = 0$.

Solve the following quadratic equations. (1 - 6)

1. $(x - 1)(x - 2) = 0$

Solution

$$(x - 1)(x - 2) = 0$$

$$x - 1 = (\quad 0 \quad) \text{ or } x - 2 = (\quad 0 \quad)$$

$$x = (\quad 1 \quad) \text{ or } x = (\quad 2 \quad)$$

2. $(x - 7)(x + 8) = 0$

Solution

$$(x - 7)(x + 8) = 0$$

$$x - 7 = 0 \text{ or } x + 8 = 0$$

$$x = \underline{7} \text{ or } x = \underline{-8}$$

3. $(4x + 1)(x + 2) = 0$

Solution

$$(4x + 1)(x + 2) = 0$$

$$4x + 1 = 0 \text{ or } x + 2 = 0$$

$$x = \underline{-\frac{1}{4}} \text{ or } x = \underline{-2}$$

4. $5x(x + 9) = 0$

Solution

$$5x(x + 9) = 0$$

$$x(x + 9) = 0$$

$$x = 0 \text{ or } x + 9 = 0$$

$$x = \underline{0} \text{ or } x = \underline{-9}$$

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Léo Pomar



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web start with 8 3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8 3 \text{ mole } \text{H}_2\text{S} \times 2 \text{ mole } \text{H}_2\text{O} / 2 \text{ mole } \text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / 1 \text{ mole } \text{H}_2\text{O} = 149 4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $137 1 \text{ g } \text{H}_2\text{O} / 149 4 \text{ g } \text{H}_2\text{O} \times 100 = 91 77$ problems 6 8 are similar to problem 5

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web percent yield is a measure of the efficiency of the experimental design yield efficiency $\text{mass of product obtained} / \text{calculated mass of the product expected} \times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160 0 \text{ g } \text{Al} \times 1 \text{ mol } \text{Al} / 27 0 \text{ g } \text{Al} \times 2 \text{ mol } \text{Al}_2\text{O}_3 / 4 \text{ mol } \text{Al} \times 101 964 \text{ g } \text{Al}_2\text{O}_3 / 302 3 \text{ g } \text{Al}_2\text{O}_3 = 26 982 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\text{yield actual} \times 100 = 260 0 \text{ g } / 260 0 \text{ g } \times 100 = 86 01$ theoretical 302 3 g

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web 7 $\text{Na}_2\text{C}_2\text{O}_4$ is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of H_2 are left over 13

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