

Half-Life

#4

If 100.0 g of carbon-14 decays until only 25.0 g of carbon is left after 11,460 y, what is the half-life of carbon-14?

1. List the given and unknown values.

Given: initial mass of sample = 100.0 g
 final mass of sample = 25.0 g
 total time of decay = 11 460 y

Unknown: number of half-lives = ? half-lives
 half-life = ? y

2. Write down the equation relating half-life, the number of half-lives, and the decay time, and rearrange it to solve for half-life.

$$\text{total time of decay} = \text{number of half-lives} \times \frac{\text{number of years}}{\text{half-life}}$$

$$\frac{\text{number of years}}{\text{half-life}} = \frac{\text{total time of decay}}{\text{number of half-lives}}$$

3. Calculate how many half-lives have passed during the decay of the 100.0 g sample.

$$\text{fraction of sample remaining} = \frac{\text{final mass of sample}}{\text{initial mass of sample}} = \frac{25.0 \text{ g}}{100.0 \text{ g}} = \frac{1}{4}$$

$$\text{after one half-life} = \frac{1}{2}; \text{ after two half-lives} = \frac{1}{2} \times \frac{1}{2} = \frac{1}{4} \text{ of sample}$$

Two half-lives have passed.

4. Solve for the half-life.

$$\frac{\text{number of years}}{\text{half-life}} = \frac{11\,460 \text{ y}}{2 \text{ half-lives}} = \frac{5730 \text{ y}}{\text{half-life}}$$

$$\text{half-life of carbon-14} = 5730 \text{ y}$$

Practice

- What is the half-life of a 100.0 g sample of nitrogen-16 that decays to 12.5 g of nitrogen-16 in 21.6 s?
- All isotopes of technetium are radioactive, but they have widely varying half-lives. If an 800.0 g sample of technetium-99 decays to 100.0 g of technetium-99 in 639 000 y, what is its half-life?
- A 208 g sample of sodium-24 decays to 13.0 g of sodium-24 within 60.0 h. What is the half-life of this radioactive isotope?

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