

DATA SHEET

Avogadro constant, N_A	$6.022 \times 10^{23} \text{ mol}^{-1}$
Volume of 1 mole ideal gas: at 100 kPa and	
at 0°C (273.15 K)	22.71 L
at 25°C (298.15 K)	24.79 L
Ionisation constant for water at 25°C (298.15 K), K_w	1.0×10^{-14}
Specific heat capacity of water	$4.18 \times 10^3 \text{ J kg}^{-1} \text{ K}^{-1}$

Some useful formulae

$$\text{pH} = -\log_{10}[\text{H}^+]$$

$$\Delta H = -m C \Delta T$$

Some standard potentials

$\text{K}^+ + \text{e}^-$	vs	K(s)	-2.94 V
$\text{Ba}^{2+} + 2\text{e}^-$	vs	Ba(s)	-2.91 V
$\text{Ca}^{2+} + 2\text{e}^-$	vs	Ca(s)	-2.87 V
$\text{Na}^+ + \text{e}^-$	vs	Na(s)	-2.71 V
$\text{Mg}^{2+} + 2\text{e}^-$	vs	Mg(s)	-2.36 V
$\text{Al}^{3+} + 3\text{e}^-$	vs	Al(s)	-1.68 V
$\text{Mn}^{2+} + 2\text{e}^-$	vs	Mn(s)	-1.18 V
$\text{H}_2\text{O} + \text{e}^-$	vs	$\frac{1}{2}\text{H}_2(\text{g}) + \text{OH}^-$	-0.83 V
$\text{Zn}^{2+} + 2\text{e}^-$	vs	Zn(s)	-0.76 V
$\text{Fe}^{2+} + 2\text{e}^-$	vs	Fe(s)	-0.44 V
$\text{Ni}^{2+} + 2\text{e}^-$	vs	Ni(s)	-0.24 V
$\text{Sn}^{2+} + 2\text{e}^-$	vs	Sn(s)	-0.14 V
$\text{Pb}^{2+} + 2\text{e}^-$	vs	Pb(s)	-0.13 V
$\text{H}^+ + \text{e}^-$	vs	$\frac{1}{2}\text{H}_2(\text{g})$	0.00 V
$\text{SO}_4^{2-} + 4\text{H}^+ + 2\text{e}^-$	vs	$\text{SO}_2(\text{aq}) + 2\text{H}_2\text{O}$	0.16 V
$\text{Cu}^{2+} + 2\text{e}^-$	vs	Cu(s)	0.34 V
$\frac{1}{2}\text{O}_2(\text{g}) + \text{H}_2\text{O} + 2\text{e}^-$	vs	2OH^-	0.40 V
$\text{Cu}^+ + \text{e}^-$	vs	Cu(s)	0.52 V
$\frac{1}{2}\text{I}_2(\text{s}) + \text{e}^-$	vs	I^-	0.54 V
$\frac{1}{2}\text{I}_2(\text{aq}) + \text{e}^-$	vs	I^-	0.62 V
$\text{Fe}^{3+} + \text{e}^-$	vs	Fe^{2+}	0.77 V
$\text{Ag}^+ + \text{e}^-$	vs	Ag(s)	0.80 V
$\frac{1}{2}\text{Br}_2(\text{l}) + \text{e}^-$	vs	Br^-	1.08 V
$\frac{1}{2}\text{Br}_2(\text{aq}) + \text{e}^-$	vs	Br^-	1.10 V
$\frac{1}{2}\text{O}_2(\text{g}) + 2\text{H}^+ + 2\text{e}^-$	vs	H_2O	1.23 V
$\frac{1}{2}\text{Cl}_2(\text{g}) + \text{e}^-$	vs	Cl^-	1.36 V
$\frac{1}{2}\text{Cr}_2\text{O}_7^{2-} + 7\text{H}^+ + 3\text{e}^-$	vs	$\text{Cr}^{3+} + \frac{7}{2}\text{H}_2\text{O}$	1.36 V
$\frac{1}{2}\text{Cl}_2(\text{aq}) + \text{e}^-$	vs	Cl^-	1.40 V
$\text{MnO}_4^- + 8\text{H}^+ + 5\text{e}^-$	vs	$\text{Mn}^{2+} + 4\text{H}_2\text{O}$	1.51 V
$\frac{1}{2}\text{F}_2(\text{g}) + \text{e}^-$	vs	F^-	2.89 V

Hsc Chemistry Data Sheet

ML Yell



Hsc Chemistry Data Sheet:

Hsc Chemistry Data Sheet Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Hsc Chemistry Data Sheet**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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