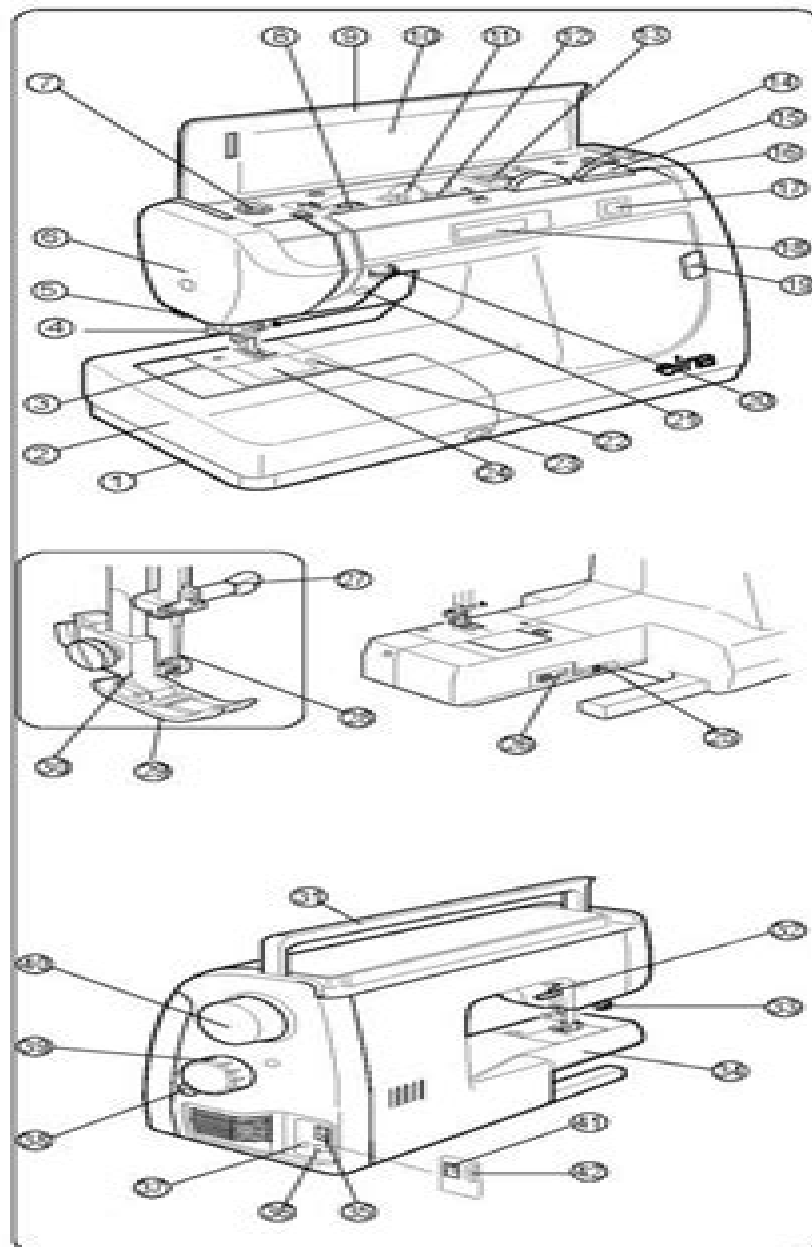




SECTION I

Machine Parts and Functions

Part Names

1. Accessory storage box
2. Extension table
3. Needle plate
4. Needle threader
5. Thread cutter
6. Face plate
7. Foot pressure dial
8. Thread tension dial
9. Top cover
10. Stitch chart
11. Large spool holder
12. Spool pin
13. Foot compartment
14. Bobbin winder spindle
15. Stitch length dial
16. Stitch width/needle position dial
17. Stitch selection display (Model 3210 only)
18. LCD screen (Model 3230 only)
19. Speed control lever (Model 3230 only)
20. Reverse stitch lever
21. Up/down needle key (Model 3230 only)
22. Bobbin cover plate release button
23. Notch -remove accessory storage box
24. Bobbin cover plate
25. Fine adjustment lever for buttonhole
26. Feed dog lever
27. Needle clamp screw
28. Needle
29. Presser foot
30. Shank
31. Carrying handle
32. Presser foot lever
33. Buttonhole lever
34. Free arm
35. On/off switch (Model 3230 only)
36. Machine socket for power (Model 3230 only)
37. Machine socket for foot control (Model 3230 only)
38. Fine adjustment screw for stretch stitches
39. Stitch selection knob
40. Flywheel (always turn towards you)
41. On/off switch (Model 3210 only)
42. Machine socket for power (Model 3210 only)

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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