Unit Test 1: Matter, Chemical Trends, & Chemical Bonding

Multiple Choice (10 marks) (KU: 3, C: 3, I: 2)

Identify the choice that best completes the statement or answers the question. Place your answers in the key below.

1.____ 2.___ 3.___ 4.___ 5.___

- The element found in the periodic table in Group 6 and Period 5 is a. tunesten. W
 c. tantalum. Ta
 - a. tungsten, w c. tantarum, ra b. molybdenum. Mo d. tellurium. Te
- 2. Which of the following groups is known for having both metallic and non-metallic properties?
 - a. alkali metals c. transition metals
 - b. alkali earth metals d. metalloids
- 3. How many electrons does a strontium atom have in its outer shell?
 - a. I c. 7 b. 2 d. 8
- 4. Which particles are responsible for the similar chemical and physical properties of elements in the same group?
 - a. protons and neutrons c. valence electrons
 - b. number of energy shells
 d. isotopes
- 5. Which of the following chemical formulas most likely represents a pure substance containing ionic bonds?
 - a. Cl₂ c. CO₂ b. SO₃ d. BaO
- Consider the equation X₁₀ + e⁻ → X⁻ + energy. The "energy" term in the equation represents
 - a. electronegativity c. ionization energy
 - b. electron affinity d. nuclear charge
- Elements A, B, C, and D (found in Groups 1–17) have atomic radii of 265 μm, 160 μm, 185 μm, and 175 μm, respectively. Which element will most likely have the lowest ionization energy?
 - a. A c. C b. B d. D
- List the following elements from smallest to largest atomic size:
 - I. barium III. strontium III. bismuth IV. iodine
 - a. III, I, II, IV b. IV, III, II, I d. IV, III, III
- 9. Which shape and bond angle is predicted by VSEPR theory for NF₃?
 - a. trigonal pyramid, less than 109.5° c. trigonal planar, 109.5°
 - b. tetrahedral, 109.5°
 d. trigonal planar, 120°

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M Mark

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