

Qualitative analysis of an organic compound showed that it contained only carbon, hydrogen and oxygen. A quantitative study of the same compound was performed, in which a 0.5096 g sample was burnt in excess oxygen to produce 0.4160 g of water and 700.7 mL of carbon dioxide, collected at 100.0 °C and 102.8 kPa.

373K

(a) Determine the empirical formula of the compound.

(10 marks)

C	H	O
0.7007 L of CO_2 $n(\text{C}) = n(\text{CO}_2)$ $PV = nRT \rightarrow n = \frac{PV}{RT}$ $n(\text{CO}_2) = \frac{102.8 \times 0.7007}{8.315 \times 373}$ $= 2.3218 \times 10^{-2} \text{ mol}$ $n(\text{C}) = nM = 0.27885 \text{ g}$	0.416 g of H_2O $n = \frac{m}{M} = \frac{0.416}{18.016} = 2.3091 \times 10^{-2} \text{ mol}$	$m(\text{O}) = 0.5096 - m(\text{C}) - m(\text{H})$ $= 0.18423 \text{ g}$ $n = \frac{m}{M} = \frac{0.18423}{16} = 1.1514 \times 10^{-2}$

EMPIRICAL FORMULA

EXAM QUESTION

$$\frac{2.3218}{1.1514} = 2$$

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Table of Contents Empirical Formula Gravimetric Questions Answers

1. Understanding the eBook Empirical Formula Gravimetric Questions Answers
 - The Rise of Digital Reading Empirical Formula Gravimetric Questions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical Formula Gravimetric Questions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical Formula Gravimetric Questions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical Formula Gravimetric Questions Answers

- Personalized Recommendations
 - Enpirical Formula Gravimetic Questions Answers User Reviews and Ratings
 - Enpirical Formula Gravimetic Questions Answers and Bestseller Lists
5. Accessing Enpirical Formula Gravimetic Questions Answers Free and Paid eBooks
 - Enpirical Formula Gravimetic Questions Answers Public Domain eBooks
 - Enpirical Formula Gravimetic Questions Answers eBook Subscription Services
 - Enpirical Formula Gravimetic Questions Answers Budget-Friendly Options
 6. Navigating Enpirical Formula Gravimetic Questions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Enpirical Formula Gravimetic Questions Answers Compatibility with Devices
 - Enpirical Formula Gravimetic Questions Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Enpirical Formula Gravimetic Questions Answers
 - Highlighting and Note-Taking Enpirical Formula Gravimetic Questions Answers
 - Interactive Elements Enpirical Formula Gravimetic Questions Answers
 8. Staying Engaged with Enpirical Formula Gravimetic Questions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Enpirical Formula Gravimetic Questions Answers
 9. Balancing eBooks and Physical Books Enpirical Formula Gravimetic Questions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Enpirical Formula Gravimetic Questions Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Enpirical Formula Gravimetic Questions Answers
 - Setting Reading Goals Enpirical Formula Gravimetic Questions Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Enpirical Formula Gravimetic Questions Answers

- Fact-Checking eBook Content of Enpirical Formula Gravimetic Questions Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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