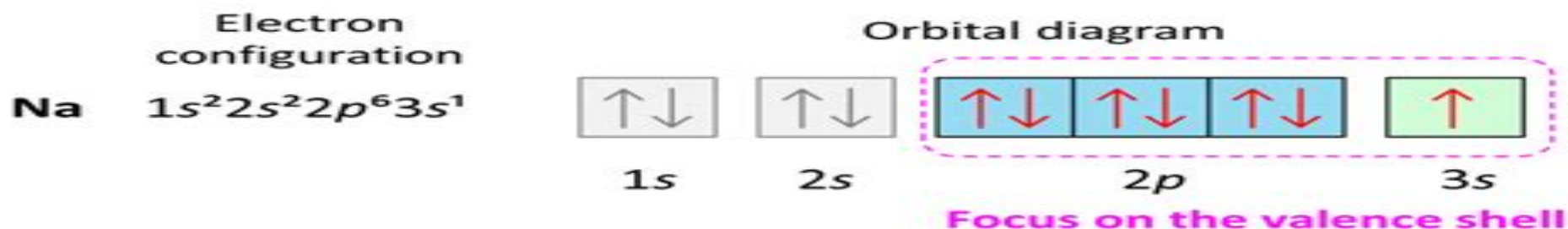


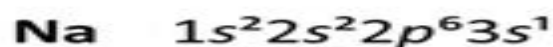
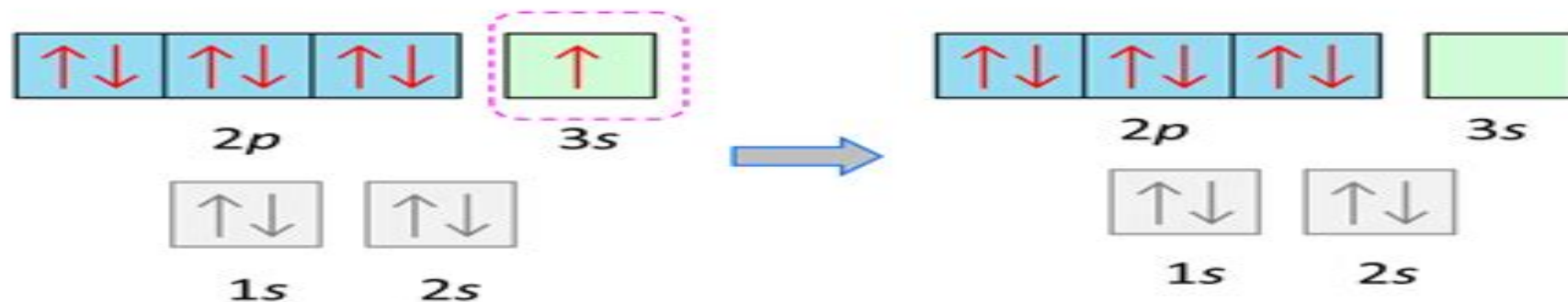
# Electron Configuration of Main Group Cations



The charge is +1, therefore, one electron is lost:



The electron must be from the outermost valence orbital –  $3s^1$



↓ Same as



Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are **isoelectronic** with the previous noble gas.

# Electrons Configurations Of Ions Lab Key

**S Baum**



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