

## Electron Configuration Practice Worksheet **KEY**

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium       $1s^2 2s^2 2p^6 3s^1$
- 2) iron       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt       $[Ar] 4s^2 3d^7$
- 7) silver       $[Kr] 5s^2 4d^9$
- 8) tellurium       $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium       $[Rn] 7s^2$
- 10) lawrencium       $[Rn] 7s^2 5f^{14} 6d^1$

Determine what elements are denoted by the following electron configurations:

- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$       **Sulfur**
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$       **Rubidium**
- 13)  $[Kr] 5s^2 4d^{10} 5p^3$       **Antimony**
- 14)  $[Xe] 6s^2 4f^{14} 5d^6$       **Osmium**
- 15)  $[Rn] 7s^2 5f^{11}$       **Einsteinium**

Determine which of the following electron configurations are not valid:

- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$       **should be  $3d^{10}$**
- 17)  $1s^2 2s^2 2p^6 3s^2 3d^5$        **$3s^2$  is impossible & "p" instead of "d"**
- 18)  $[Ra] 7s^2 5f^8$       **Ra is not a noble gas**
- 19)  $[Kr] 5s^2 4d^{10} 5p^5$       **Valid**
- 20)  $[Xe]$       **Element can't be its own electron configuration.**

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**F Rizvi**



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## **Table of Contents Electrons Configuration Worksheet And Lots More Answers**

1. Understanding the eBook Electrons Configuration Worksheet And Lots More Answers
  - The Rise of Digital Reading Electrons Configuration Worksheet And Lots More Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Electrons Configuration Worksheet And Lots More Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electrons Configuration Worksheet And Lots More Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electrons Configuration Worksheet And Lots More Answers
  - Personalized Recommendations
  - Electrons Configuration Worksheet And Lots More Answers User Reviews and Ratings
  - Electrons Configuration Worksheet And Lots More Answers and Bestseller Lists
5. Accessing Electrons Configuration Worksheet And Lots More Answers Free and Paid eBooks
  - Electrons Configuration Worksheet And Lots More Answers Public Domain eBooks
  - Electrons Configuration Worksheet And Lots More Answers eBook Subscription Services
  - Electrons Configuration Worksheet And Lots More Answers Budget-Friendly Options

6. Navigating Electrons Configuration Worksheet And Lots More Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Electrons Configuration Worksheet And Lots More Answers Compatibility with Devices
  - Electrons Configuration Worksheet And Lots More Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electrons Configuration Worksheet And Lots More Answers
  - Highlighting and Note-Taking Electrons Configuration Worksheet And Lots More Answers
  - Interactive Elements Electrons Configuration Worksheet And Lots More Answers
8. Staying Engaged with Electrons Configuration Worksheet And Lots More Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electrons Configuration Worksheet And Lots More Answers
9. Balancing eBooks and Physical Books Electrons Configuration Worksheet And Lots More Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electrons Configuration Worksheet And Lots More Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Electrons Configuration Worksheet And Lots More Answers
  - Setting Reading Goals Electrons Configuration Worksheet And Lots More Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electrons Configuration Worksheet And Lots More Answers
  - Fact-Checking eBook Content of Electrons Configuration Worksheet And Lots More Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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