

# Electron Configuration of Main Group Cations

Electron configuration

**Na**  $1s^2 2s^2 2p^6 3s^1$

Orbital diagram

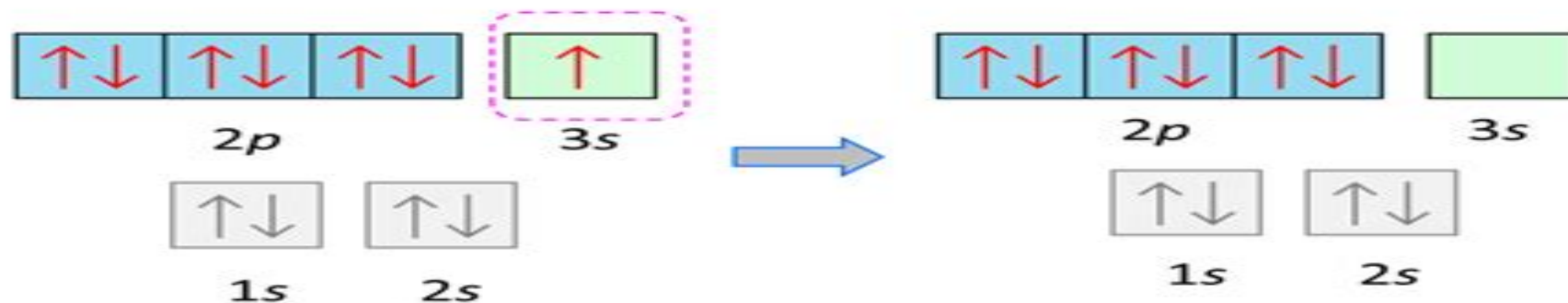


Focus on the valence shell

The charge is +1, therefore, one electron is lost:



The electron must be from the outermost valence orbital –  $3s^1$



**Na**  $1s^2 2s^2 2p^6 3s^1$

**Na<sup>+</sup>**  $1s^2 2s^2 2p^6$

↓ Same as

**Ne**  $1s^2 2s^2 2p^6$



ChemistrySteps.com

Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are **isoelectronic** with the previous noble gas.

# Electron Configurations Of Atoms And Ions Small Scale Lab

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