

In the space below, write the unabbreviated electron configurations of the following elements:

Drag the correct electron configuration to the correct element below.

8  
**O**  
oxygen  
16.0

11  
**Na**  
sodium  
23.0

26  
**Fe**  
iron  
55.8

35  
**Br**  
bromine  
79.9

$1s^2 2s^2 2p^6 3s^1$

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$

$1s^2 2s^2 2p^4$

# Electron Configuration Worksheet Packet And Lots More Answers

**Yijin Wang**



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web the balanced chemical equation for this reaction is  $\text{mg} + 2\text{hno}_3 \text{aq} \rightarrow \text{mg}(\text{no}_3)_2 \text{aq} + \text{h}_2 \text{g}$  in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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produce water and calcium nitrate

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you balance magnesium and nitric acid you ll need to count all of atoms on both directions of the chemical equation to  
balance mg hno3 mm no3 2 no2 h2o

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with an eye toward strengthening the content available in owlv2 end of chapter questions and updating the presentation  
nomenclature changes and the

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chemical reaction for this reaction we have a single displacement reaction balancing strategies in this single displacement  
reaction the mg replaces the h in the hno3

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mg 2 2 oh 1 2 h 1 2no 3 1 2h 2 o mg 2 2 no 3 1 net ionic equation 2 oh 1 2 h 1 2h 2 o sulfuric acid and lithium hydroxide  
balanced total equation h 2 so 4 aq 2 lih aq 2 h 2