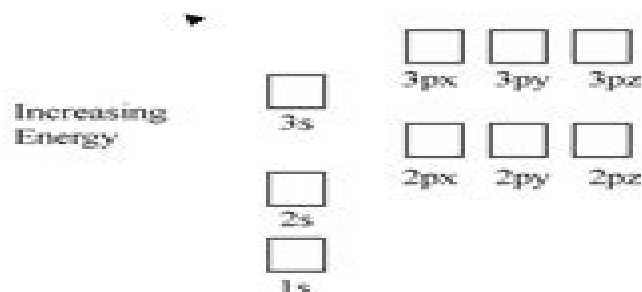


## Electron configuration

The area in which there is 90% probability of finding an electron is called an orbital. An orbital can hold two electrons.

These orbitals can be represented in the following way:

### Aufbau Diagram



Each box represents an orbital. Electrons are designated by arrows pointing up or down which represent clockwise or counterclockwise spins.




To determine where the electrons of a particular element are located you use the following rules:

**RULE 1: An electron will always be found in the lowest energy orbital available that has room.**

Therefore the electron configuration of Hydrogen is:  This is designated as:  $1s^1$

**RULE 2: An orbital may hold a maximum of 2 electrons.**

**RULE 3: The Pauli Exclusion Principle – Two electrons sharing an orbital will have opposite spins.**

Therefore the electron configuration of Helium is:  This is designated as:  $1s^2$

1. Draw an orbital diagram and write the electron configuration for Lithium, Beryllium, and Boron.

Li  $1s^2 2s^1$

Be  $1s^2 2s^2$

B  $1s^2 2s^2 2p^1$

Orbital diagrams for Li, Be, and B are shown below. For Li, the 1s orbital is filled with two electrons (up and down arrows), and the 2s orbital has one electron (up arrow). For Be, the 1s orbital is filled with two electrons, and the 2s orbital is also filled with two electrons. For B, the 1s orbital is filled with two electrons, the 2s orbital is filled with two electrons, and the 2p orbital has one electron (up arrow).

# Electron Configuration Packet Key

**Mark Weller, Mark T. Weller, Tina  
Overton, Jonathan Rourke, Fraser  
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