

Electron Configuration Lab

Name _____ KEY _____ Period _____ Date _____

Directions: To understand that electrons fill up the LOWEST energy orbitals first, but that they spread out as much as possible in orbitals.

Part 1: Questions

1. Write the Aufbau Principle.

Electrons will fill the lowest energy orbitals first.

2. Which "rule" or "principle" states that electrons spread out as much as possible within a sublevel before they pair up in the same orbital? **Hund's Rule**

3. In this lab, one cup represents an orbital in the atom. How many cups would be needed to represent the "2p" sublevel?

three (3)

4. How many cups would be needed to represent an atom that has the 1s, 2s, 2p, 3s, 3p sublevels _____?

ten (10)

5. In this lab, beans represent the electrons. How many beans can go into a single cup?

two (2)

Part 2:

Part 2a

Electron Configuration	Total # of Electrons	Element Name	Teacher's Initials
$1s^2 2s^2 2p^3 3s^2 3p^3$	15	Phosphorus	

Part 2b

Station	Electron Configuration	Total # of Electrons
Station 1	$1s^2 2s^2 2p^3$	10
Station 2	$1s^2 2s^2 2p^4$	8
Station 3	$1s^2 2s^2 2p^6 3s^2 3p^2 4s^2$	20
Station 4	$1s^2 2s^2 2p^6 3s^2 3p^3$	18
Station 5	$1s^2 2s^2 2p^3$	9
Station 6	$1s^2 2s^2$	4
Station 7	$1s^2 2s^2 2p^3$	8
Station 8	$1s^2 2s^2$	4
Station 9	$1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$	18
Station 10	$1s^2 2s^2 2p^3$	9

Part 2c: Questions

1. Using a Periodic Table, identify the elements at each station and complete the table to the right.

2. Which orbital is lower in energy -- 4s or 3d?

4s

3. Write the electron configuration for Zinc.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10}$

4. Write the electron configuration for Xenon.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6$

5. Write the electron configuration for Bismuth.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^3$

Station	Element Name
Station 1	Neon (Ne)
Station 2	Oxygen (O)
Station 3	Calcium (Ca)
Station 4	Argon (Ar)
Station 5	Boron (B)
Station 6	Beryllium (Be)
Station 7	Carbon (C)
Station 8	Lithium (Li)
Station 9	Potassium (K)
Station 10	Nitrogen (N)

Electron Configuration Lab

Siddappa N.Byrareddy



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Table of Contents Electron Configuration Lab

1. Understanding the eBook Electron Configuration Lab
 - The Rise of Digital Reading Electron Configuration Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Lab
 - Personalized Recommendations
 - Electron Configuration Lab User Reviews and Ratings
 - Electron Configuration Lab and Bestseller Lists
5. Accessing Electron Configuration Lab Free and Paid eBooks
 - Electron Configuration Lab Public Domain eBooks
 - Electron Configuration Lab eBook Subscription Services
 - Electron Configuration Lab Budget-Friendly Options

6. Navigating Electron Configuration Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Lab Compatibility with Devices
 - Electron Configuration Lab Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Lab
 - Highlighting and Note-Taking Electron Configuration Lab
 - Interactive Elements Electron Configuration Lab
8. Staying Engaged with Electron Configuration Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Lab
9. Balancing eBooks and Physical Books Electron Configuration Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Lab
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration Lab
 - Setting Reading Goals Electron Configuration Lab
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration Lab
 - Fact-Checking eBook Content of Electron Configuration Lab
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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