

Chapter 2: Atoms, Molecules and Ions

Dalton's Atomic Theory (1805)

1. All matter consists of **atoms**; tiny indivisible particles of an element that cannot be created or destroyed.
2. Atoms of one element **cannot** be subdivided, created, or destroyed during chemical reactions. (wrong)
3. Atoms of an element are **identical** in mass and other properties and are different from the atoms of any other element. (wrong)
4. Compounds result from the chemical combination of a **specific ratio** of atoms of different elements.

Atomic Structure: Electrons

- J.J. Thomson (1897) Cathode Ray Tube Experiment
- Observations:
 - Ray bends in magnetic field
 - Ray bends toward positive plate in electric field
 - Ray is identical for any cathode
- Conclusions:
 - Ray consists of charged particles
 - Ray consists of negative particles
 - These particles are found in ALL matter

Cathode Ray Tube experiment using magnets led to the discovery of electrons

Thomson's Model of the Atom

- Plum-pudding model:
 - E- distributed throughout diffuse, positively charged "squishy" sphere



Atomic Structure: The Nucleus

- Rutherford's Experiment
 - Bombard a thin gold foil with positively charged particles to test Thomson's model of the atom
 - Observations:
 - Atom is made up of smaller subatomic particles
 - Protons
 - Neutrons
 - Electrons (mass of only ab 1/2000th of other particles)
 - Conclusions

Genral Chemistry Study Guide

Darrell Ebbing, Steven D. Gammon



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Table of Contents Genral Chemistry Study Guide

1. Understanding the eBook Genral Chemistry Study Guide
 - The Rise of Digital Reading Genral Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Genral Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Genral Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Genral Chemistry Study Guide
 - Personalized Recommendations
 - Genral Chemistry Study Guide User Reviews and Ratings
 - Genral Chemistry Study Guide and Bestseller Lists
5. Accessing Genral Chemistry Study Guide Free and Paid eBooks
 - Genral Chemistry Study Guide Public Domain eBooks
 - Genral Chemistry Study Guide eBook Subscription Services
 - Genral Chemistry Study Guide Budget-Friendly Options

6. Navigating Genral Chemistry Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Genral Chemistry Study Guide Compatibility with Devices
 - Genral Chemistry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Genral Chemistry Study Guide
 - Highlighting and Note-Taking Genral Chemistry Study Guide
 - Interactive Elements Genral Chemistry Study Guide
8. Staying Engaged with Genral Chemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Genral Chemistry Study Guide
9. Balancing eBooks and Physical Books Genral Chemistry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Genral Chemistry Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Genral Chemistry Study Guide
 - Setting Reading Goals Genral Chemistry Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Genral Chemistry Study Guide
 - Fact-Checking eBook Content of Genral Chemistry Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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