

The isotope fractionation factor

The isotope fractionation factor is defined as: $\alpha_b^a = \frac{R_a}{R_b}$

where R_A , R_B are the isotope ratios in two phases (e.g., carbonate and water, or water vapor and water)

For example, consider: $\text{H}_2\text{O}(l) \leftrightarrow \text{H}_2\text{O}(v)$ at 25°C

$$\alpha_v^l(O) = \frac{R_l}{R_v} = \frac{\left(^{18}\text{O}/^{16}\text{O}\right)_l}{\left(^{18}\text{O}/^{16}\text{O}\right)_v} = 1.0092$$

The differences in isotope ratios are relatively small and are expressed as parts per thousand (per mil) deviations from a standard.

Fractionation Formula Examples

Raffaella Di Napoli

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Fractionation Formula Examples

1. Understanding the eBook Fractionation Formula Examples
 - The Rise of Digital Reading Fractionation Formula Examples
 - Advantages of eBooks Over Traditional Books
2. Identifying Fractionation Formula Examples
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Fractionation Formula Examples
 - User-Friendly Interface
4. Exploring eBook Recommendations from Fractionation Formula Examples
 - Personalized Recommendations
 - Fractionation Formula Examples User Reviews and Ratings
 - Fractionation Formula Examples and Bestseller Lists

5. Accessing Fractionation Formula Examples Free and Paid eBooks
 - Fractionation Formula Examples Public Domain eBooks
 - Fractionation Formula Examples eBook Subscription Services
 - Fractionation Formula Examples Budget-Friendly Options
6. Navigating Fractionation Formula Examples eBook Formats
 - ePub, PDF, MOBI, and More
 - Fractionation Formula Examples Compatibility with Devices
 - Fractionation Formula Examples Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Fractionation Formula Examples
 - Highlighting and Note-Taking Fractionation Formula Examples
 - Interactive Elements Fractionation Formula Examples
8. Staying Engaged with Fractionation Formula Examples
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Fractionation Formula Examples
9. Balancing eBooks and Physical Books Fractionation Formula Examples
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Fractionation Formula Examples
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Fractionation Formula Examples
 - Setting Reading Goals Fractionation Formula Examples
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Fractionation Formula Examples
 - Fact-Checking eBook Content of Fractionation Formula Examples
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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