Section 5.1: Models of the Atom

How many sublevels are in the following principal energy levels? Which sublevels?

a. n = 1 One - s

e. n = 5 Four -s, p. d. f

b. n=2 Two $-s_n p$

f. n = 6 Three - s, p, d

c. n = 3 Three -s, p. d

g. n = 7 Two - s, p

- d. n = 4 Four s, p, d, f
- 2. How many orbitals are in the following sublevels?

a. Is sublevel One

d. 4f sublevel Seven

b. 2p sublevel Three

e. 5s sublevel One

c. 3d sublevel Five

- f. 6p sublevel Three
- How many electrons can each of the following sublevels hold (maximum)?

a. s Two

c. d Ten

b. p Six

d. f Fourteen

- 4. What is the shape of an s sublevel orbital? Sphere
- What is the shape of a p sublevel orbital? Dumbbell
- What was the problem with Rutherford's planetary model of the atom? It could not explain chemical properties
- 7. Bohr proposed that electrons were found in specific circular paths called orbits
- What happens to the distance between energy levels as you increase in energy? It decreases.
- What does "electron cloud" mean? A region of space with the highest probability, or likelihood, of finding an electron. Not an exact path.
- 10. How do electrons move from energy level to energy level? Gain or lose energy

Section 5.2: Electron Arrangement in Atoms

- Electrons arrange themselves around the nucleus in order to make the most stable arrangement.
- 2. What are the three rules for writing electron configurations? Give the definition of each.
 - Aufbau's Principle electrons occupy the lowest energy levels first
 - b. Pauli Exclusion Principle electrons in the same orbital have opposite spin
 - c. Hund's Rule electrons enter orbitals individually before pairing up

Electrons In Atoms Chapter Test A Worksheet Answers

KJ Lindholm-Leary

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