

# 5

## ELECTRONS IN ATOMS

### Practice Problems

In your notebook, solve the following problems.

#### SECTION 5.1 MODELS OF THE ATOM

- How many sublevels are in the following principal energy levels?
  - $n = 1$
  - $n = 2$
  - $n = 3$
  - $n = 4$
  - $n = 5$
  - $n = 6$
- How many orbitals are in the following sublevels?
  - 1s sublevel
  - 5s sublevel
  - 4d sublevel
  - 4f sublevel
  - 7s sublevel
  - 3p sublevel
  - fifth principal energy level
  - 6d sublevel
- What are the types of sublevels and number of orbitals in the following energy levels?
  - $n = 1$
  - $n = 2$
  - $n = 3$
  - $n = 4$
  - $n = 5$

#### SECTION 5.2 ELECTRON ARRANGEMENT IN ATOMS

- Write a complete electron configuration of each atom.
  - hydrogen
  - vanadium
  - magnesium
  - barium
  - bromine
  - sulfur
  - krypton
  - arsenic
  - radon

#### SECTION 5.3 PHYSICS AND THE QUANTUM MECHANICAL MODEL

- What is the wavelength of the radiation whose frequency is  $5.00 \times 10^{15} \text{ s}^{-1}$ ? In what region of the electromagnetic spectrum is this radiation?
- An inexpensive laser that is available to the public emits light that has a wavelength of 670 nm. What are the color and frequency of the radiation?
- What is the energy of a photon whose frequency is  $2.22 \times 10^{14} \text{ s}^{-1}$ ?
- What is the frequency of a photon whose energy is  $6.00 \times 10^{-19} \text{ J}$ ?
- Arrange the following types of electromagnetic radiation in order of increasing frequency.
  - infrared
  - cosmic rays
  - visible light
  - radio waves
  - microwaves
  - ultraviolet
- Suppose that your favorite AM radio station broadcasts at a frequency of 1600 kHz. What is the wavelength in meters of the radiation from the station?

# Electrons In Atoms Answer Key

**Mark Steyvers**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents **Electrons In Atoms Answer Key****

1. Understanding the eBook **Electrons In Atoms Answer Key**
  - The Rise of Digital Reading **Electrons In Atoms Answer Key**
  - Advantages of eBooks Over Traditional Books
2. Identifying **Electrons In Atoms Answer Key**
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an **Electrons In Atoms Answer Key**
  - User-Friendly Interface
4. Exploring eBook Recommendations from **Electrons In Atoms Answer Key**
  - Personalized Recommendations
  - **Electrons In Atoms Answer Key** User Reviews and Ratings

- Electrons In Atoms Answer Key and Bestseller Lists
- 5. Accessing Electrons In Atoms Answer Key Free and Paid eBooks
  - Electrons In Atoms Answer Key Public Domain eBooks
  - Electrons In Atoms Answer Key eBook Subscription Services
  - Electrons In Atoms Answer Key Budget-Friendly Options
- 6. Navigating Electrons In Atoms Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Electrons In Atoms Answer Key Compatibility with Devices
  - Electrons In Atoms Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electrons In Atoms Answer Key
  - Highlighting and Note-Taking Electrons In Atoms Answer Key
  - Interactive Elements Electrons In Atoms Answer Key
- 8. Staying Engaged with Electrons In Atoms Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electrons In Atoms Answer Key
- 9. Balancing eBooks and Physical Books Electrons In Atoms Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electrons In Atoms Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Electrons In Atoms Answer Key
  - Setting Reading Goals Electrons In Atoms Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electrons In Atoms Answer Key
  - Fact-Checking eBook Content of Electrons In Atoms Answer Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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