

# 8 • Electron Configurations & Periodicity

## WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e <sup>-</sup> s: <u>18</u>	Element: Mg # of e <sup>-</sup> s: <u>12</u>	Element: N # of e <sup>-</sup> s: <u>7</u>	Element: Li # of e <sup>-</sup> s: <u>3</u>	Element: P # of e <sup>-</sup> s: <u>15</u>	Element: Cl # of e <sup>-</sup> s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar  $1s^2 2s^2 2p^6 3s^2 3p^6$   
 Ar  $[\text{Ne}] 3s^2 3p^6$
- Mg  $1s^2 2s^2 2p^6 3s^2$   
 Mg  $[\text{Ne}] 3s^2$
- N  $1s^2 2s^2 2p^3$   
 N  $[\text{He}] 2s^2 2p^3$
- Li  $1s^2 2s^1$   
 Li  $[\text{He}] 2s^1$
- P  $1s^2 2s^2 2p^6 3s^2 3p^3$   
 P  $[\text{Ne}] 3s^2 3p^3$
- Cl  $1s^2 2s^2 2p^6 3s^2 3p^5$   
 Cl  $[\text{Ne}] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



# Electron Configuration Review Answer Sheet

**United States. National Bureau of  
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## **Table of Contents Electron Configuration Review Answer Sheet**

1. Understanding the eBook Electron Configuration Review Answer Sheet
  - The Rise of Digital Reading Electron Configuration Review Answer Sheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Review Answer Sheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electron Configuration Review Answer Sheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Review Answer Sheet
  - Personalized Recommendations
  - Electron Configuration Review Answer Sheet User Reviews and Ratings
  - Electron Configuration Review Answer Sheet and Bestseller Lists

5. Accessing Electron Configuration Review Answer Sheet Free and Paid eBooks
  - Electron Configuration Review Answer Sheet Public Domain eBooks
  - Electron Configuration Review Answer Sheet eBook Subscription Services
  - Electron Configuration Review Answer Sheet Budget-Friendly Options
6. Navigating Electron Configuration Review Answer Sheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Electron Configuration Review Answer Sheet Compatibility with Devices
  - Electron Configuration Review Answer Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electron Configuration Review Answer Sheet
  - Highlighting and Note-Taking Electron Configuration Review Answer Sheet
  - Interactive Elements Electron Configuration Review Answer Sheet
8. Staying Engaged with Electron Configuration Review Answer Sheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electron Configuration Review Answer Sheet
9. Balancing eBooks and Physical Books Electron Configuration Review Answer Sheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electron Configuration Review Answer Sheet
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration Review Answer Sheet
  - Setting Reading Goals Electron Configuration Review Answer Sheet
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration Review Answer Sheet
  - Fact-Checking eBook Content of Electron Configuration Review Answer Sheet
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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