

**M** A buffer solution using the conjugate acid-base pair acetic acid and acetate ions. Write chemical equations showing the reactions that take place when  $H^+$  and when  $OH^-$  are added to the buffer.

4. Acetic acid has a  $K_a$  value of  $1.8 \times 10^{-5}$ . How many grams of  $\text{NaC}_2\text{H}_3\text{O}_2$  would have to be added to 100 ml. of 0.10 M  $\text{HC}_2\text{H}_3\text{O}_2$  to prepare a buffer with a pH of 4.50?

$$(0.057M)(.100L) = 0.0057\text{mol}$$
$$= 0.057 M$$

Acetic acid,  $\text{HC}_2\text{H}_3\text{O}_2$ , 0.1 M, 20 ml.

Sodium chloride, NaCl, 0.1 M, 40 ml.

Ammonia, NH<sub>3</sub>, 0.1 M, 20 ml.

Sodium Hydroxide, NaOH, 0.1 M, 5 ml.

Ammonium chloride,  $\text{NH}_4\text{Cl}$ , 0.1 M, 20 ml.

Sodium hydroxide, NaOH, 0.2 M, 30 ml.

Hydrochloric acid, HCl, 0.1 M, 5 ml.

Unknown solid acid anhydride, 2.0 g

Phenylhydrazine, 1.0 g, 1 ml.

Water, distilled or deionized

Sodium acetate,  $\text{NaC}_2\text{H}_3\text{O}_2$ , 0.1 M, 20 ml.

$$[\text{C}_2\text{H}_3\text{O}_2^-] = \frac{K_a [\text{HC}_2\text{H}_3\text{O}_2]}{[\text{H}^+]} = \frac{(1.8 \times 10^{-5})(0.10)}{3.16 \times 10^{-5}}$$

Hickman, C. C. 1993.

2

Penetration, 1992 and 1993

CLINTON, N.Y.

Abstract-type reports 3

Buret, ring stand, and beuret clapper

Wash bottle with distilled or deionized water.

Reference: 150 mL, 1

# Flinn Buffer Solutions Lab Answer Key

**Robert R. Redfield**



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