

*Honors Chemistry Midterm: Review Packet***PART 1**

1. Write in the number of significant figures for each of the following measurements:

3 3.00 mL 3 0.00290 g 4 50.00 m 2 0.070 kg 1 400 L

2. Using the following masses: 0.0562 kg, 124.213 g and 1635 mg.
- ← convert to grams*
-
- The answer should be reported as:
-
- g

56.2g 124.213g 1.635g

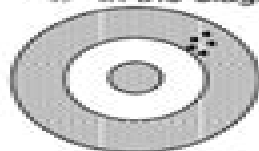
3. a) A 5.75 mL sample of mercury has a measured mass of 77.05 g. The density is
- 13.49/mL

$$D = \frac{m}{V} = \frac{77.05g}{5.75mL} =$$

- b) Mercury's accepted density is 13.53 g/mL. The % error in your measurement is:
- .9617%

$$\frac{|\text{actual} - \text{exp}|}{\text{actual}} \times 100 \quad \frac{|13.53g/mL - 13.49g/mL|}{13.53g/mL} \times 100$$

4. In the diagram below:



Is this person accurate? no - not in center

Is this person precise? yes - close together

5. Identify the following as a chemical change (CC), physical change (PC), chemical property (CP), or physical property (PP):

a) burning wood CC b) mass of wood PP

c) flammability of wood CP d) chopping up wood PC

Honors Chemistry Semester Review Packet Answers

JR Anderson



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Table of Contents Honors Chemistry Semester Review Packet Answers

1. Understanding the eBook Honors Chemistry Semester Review Packet Answers
 - The Rise of Digital Reading Honors Chemistry Semester Review Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Honors Chemistry Semester Review Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Honors Chemistry Semester Review Packet Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Honors Chemistry Semester Review Packet Answers
 - Personalized Recommendations
 - Honors Chemistry Semester Review Packet Answers User Reviews and Ratings
 - Honors Chemistry Semester Review Packet Answers and Bestseller Lists
5. Accessing Honors Chemistry Semester Review Packet Answers Free and Paid eBooks
 - Honors Chemistry Semester Review Packet Answers Public Domain eBooks
 - Honors Chemistry Semester Review Packet Answers eBook Subscription Services
 - Honors Chemistry Semester Review Packet Answers Budget-Friendly Options
6. Navigating Honors Chemistry Semester Review Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Honors Chemistry Semester Review Packet Answers Compatibility with Devices
 - Honors Chemistry Semester Review Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Honors Chemistry Semester Review Packet Answers
 - Highlighting and Note-Taking Honors Chemistry Semester Review Packet Answers
 - Interactive Elements Honors Chemistry Semester Review Packet Answers
8. Staying Engaged with Honors Chemistry Semester Review Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Honors Chemistry Semester Review Packet Answers
9. Balancing eBooks and Physical Books Honors Chemistry Semester Review Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Honors Chemistry Semester Review Packet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Honors Chemistry Semester Review Packet Answers
 - Setting Reading Goals Honors Chemistry Semester Review Packet Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Honors Chemistry Semester Review Packet Answers
 - Fact-Checking eBook Content of Honors Chemistry Semester Review Packet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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