

Balancing Charges on Ions

Atoms of different elements combine with one another to form compounds. The empirical formula of an ionic compound indicates the kinds of atoms that are present in the compound, as well as the relative number (ratio) of each kind of atom. Let's investigate how the formula of an ionic compound can be determined experimentally.

- Toxic composition
- Polymorphic forms
- Empirical formula
- Physicochemical reactions

An ionic compound is composed of ions—atoms or groups of atoms that have a positive or negative charge. Oppositely charged ions arrange themselves in an ordered, three-dimensional structure called a crystal lattice. The net attractive forces among oppositely charged ions in the crystal structure are called ionic bonds. Although composed of charged ions, ionic compounds are electrically neutral. The ratio of oppositely charged ions in the crystal structure is such that the positive charge contributed by the cations is equal in magnitude to the negative charge contributed by the anions. There is no net or overall change in an ionic compound.

The empirical formulae of an ionic compound indicates the smallest whole-number ratio of each type of ion in the crystal structure and is called a formula unit. For example, magnesium chloride has the empirical formula MgCl_2 . Magnesium cations (Mg^{2+}) and chloride anions (Cl^{-}) combine in a 1:2 ratio to form the MgCl_2 formula unit. The overall charge on ionic compounds is always zero.

Some ions consist of a charged group of covalently bonded atoms. Such ions are called *polyatomic ions*. An example is the nitrate ion (NO_3^-), which contains one nitrogen atom and three oxygen atoms, and has an overall charge of -1 . In calcium nitrate, calcium (Ca^{2+}), ions combine with nitrate ions in a 1:2 ratio in order to balance the positive and negative charges. The empirical formula for calcium nitrate is $\text{Ca}(\text{NO}_3)_2$. Parentheses are used around the nitrate ions to show that the subscript "2" pertains to the nitrate ion as a whole.

Many ionic compounds can be prepared in the lab using precipitation reactions. When solutions of two ionic compounds are combined, the ions may rearrange to form a new ionic compound that is insoluble in water. An example of this type of reaction is the formation of solid barium sulfate when barium chloride and sodium sulfate are combined in solution (Equation 1). In Equation 1, only the ions that form the precipitate are represented. This makes it easier to recognize what happens in the precipitation reaction.



According to the balanced equation for this reaction, barium ions (Ba^{2+}) combine with sulfate ions (SO_4^{2-}) in a 1:1 ratio to form barium sulfate (BaSO_4). This ratio can be observed experimentally by the fully mixing BaCl_2 and Na_2SO_4 solutions containing equal amounts (concentrations) of barium and sulfate ions, respectively. The maximum amount of precipitate will be obtained, where equal volumes (a 1:1 ratio) of the two solutions are combined. A similar approach can also be used to determine the formula of an unknown ionic compound.

The purpose of this experiment is to determine the stoichiometric formula of an unknown ionic compound. Two solutions containing equal amounts (concentration) of two reactants ions will be combined in a series of reactions. In each reaction, the total volume of the two solutions will be held constant, while the volume ratio of the reactants is varied. The amount of precipitate obtained in each reaction will be measured and plotted against the volume ratio to find the stoichiometric formula of the product.

Flinn Chemtopic Labs Ionic Compound Answers

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