

## TEMPERATURE

W

ith this chapter we begin our study of thermal physics, the branch of physics that deals with the changes in the properties of systems that occur when work is done on (or by) them and heat energy is added to (or taken from) them. For systems such as confined gases the properties involved are their pressure, volume, temperature, energy, and—as we will come to learn—entropy, a property that we will introduce in Chapter 24.

This chapter deals with temperature, a concept that deeply underlies all of the laws of thermodynamics. We have used this concept in earlier chapters; now we must define it precisely, as we have done for all other physical concepts that we have encountered. We also introduce the concept of an ideal gas, which will serve as a convenient system to use in analyzing and illustrating the laws of thermodynamics.

## 21-1 TEMPERATURE AND THERMAL EQUILIBRIUM

We all have an ingrained sense of temperature and, indeed, we have used this concept freely in earlier chapters. In this chapter we wish to define temperature in a rigorous way. Just as we went beyond our sense of “push” and “pull” in defining force, we need to go beyond our sense of “hot” and “cold” in defining temperature. Before we can deal directly with temperature, however, we must first establish the concept of *thermal equilibrium*, which is concerned with the question of whether or not the temperatures of two systems are equal.

Figure 21-1a shows two systems *A* and *B*, which, among many possibilities, might be blocks of metal or confined gases. They are isolated from one another and from their environment, by which we mean that neither energy nor matter can enter or leave either system. For example, the systems may be surrounded by walls made of thick slabs of Styrofoam, presumed to be both rigid and impermeable. Such walls are said to be *adiabatic*, which you can think of as meaning *thermally insulating*. Changes in the measured properties of either system have no effect on the properties of the other system.

As Fig. 21-1b shows, we can replace the adiabatic wall that separates the two systems with one that permits the

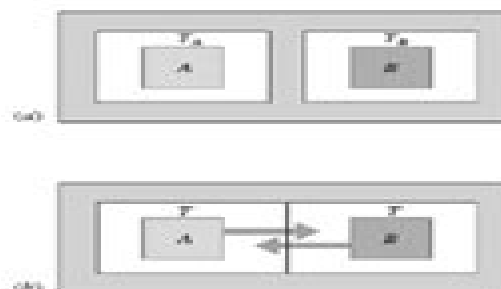


FIGURE 21-1. (a) Systems *A* and *B* are separated by an adiabatic wall. The systems have different temperatures  $T_A$  and  $T_B$ . (b) Systems *A* and *B* are separated by a diathermic wall, which permits energy to be exchanged between the systems. The systems will eventually come to thermal equilibrium, upon which they have the same temperature  $T$ .

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