

CHM 2210 Lab: Melting Point

Introduction

The melting point of a solid is the temperature at which the solid melts (changes to a liquid) at 1 atm of pressure. Melting point is determined by slow heating of a small solid sample placed in a capillary tube. The melting point should be reported as a melting point range: from the temperature at which the first droplet of liquid is observed to the temperature at which the sample becomes a clear liquid, for example: 96.7 – 98.9°C.

A pure solid (pure enough for most purposes) usually has a sharp melting point which means that it melts within a range of 2°C or less.

An impure solid melts at a broader range (3°C or more) and at depressed temperature.

Using melting point to find the identity of a solid is not sufficient because there are too many substances with the same melting points.

However, you can use a mixed melting point as a tool for identification of a solid.

If you mix two different solids with the same melting point, the melting point of the mixture will be broad and depressed. The melting point of a solid and a mixture of this solid with the authentic solid should be the same.

Example:

Measured/Observed melting point of an unknown solid X: 128.0 – 129.5°C

Melting point of known solid A: 130°C

Melting point of known solid B: 158°C

Measured melting point of mixture X + A: 127.5 – 129.0°C

Measured melting point of mixture X + B: 133.5 – 142.0°C

Identity of unknown X: solid A

In the future you will measure melting points of solids mostly to determine their purity.

- **Preparation of Sample**

A solid sample should be dried and crushed with a metal spatula or bottom of a test tube.

If a solid is used for mixed melting point, the 1:1 mixture should be crushed together into a homogeneous mixture.

The pulverized solid should be mounded up and a small (1 – 2 mm) plug of sample should be pushed into the open end of the capillary tube. The sample should be knocked down to the sealed end of the capillary tube by tapping its sealed end on the surface or dropping it into a long glass tubing.

The sample should be 1-2 mm long and packed firmly.

If the sample is too large, the measured melting point may be too high.

If the sample is loosely packed, it may be difficult to determine the melting point range because of the air present.

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Table of Contents Melting Point Lab For High School

1. Understanding the eBook Melting Point Lab For High School
 - The Rise of Digital Reading Melting Point Lab For High School
 - Advantages of eBooks Over Traditional Books
2. Identifying Melting Point Lab For High School
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Melting Point Lab For High School
 - User-Friendly Interface
4. Exploring eBook Recommendations from Melting Point Lab For High School
 - Personalized Recommendations
 - Melting Point Lab For High School User Reviews and Ratings

- Melting Point Lab For High School and Bestseller Lists
- 5. Accessing Melting Point Lab For High School Free and Paid eBooks
 - Melting Point Lab For High School Public Domain eBooks
 - Melting Point Lab For High School eBook Subscription Services
 - Melting Point Lab For High School Budget-Friendly Options
- 6. Navigating Melting Point Lab For High School eBook Formats
 - ePub, PDF, MOBI, and More
 - Melting Point Lab For High School Compatibility with Devices
 - Melting Point Lab For High School Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Melting Point Lab For High School
 - Highlighting and Note-Taking Melting Point Lab For High School
 - Interactive Elements Melting Point Lab For High School
- 8. Staying Engaged with Melting Point Lab For High School
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Melting Point Lab For High School
- 9. Balancing eBooks and Physical Books Melting Point Lab For High School
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Melting Point Lab For High School
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Melting Point Lab For High School
 - Setting Reading Goals Melting Point Lab For High School
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Melting Point Lab For High School
 - Fact-Checking eBook Content of Melting Point Lab For High School
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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