

### Chapter 3: Periodic Trends (NB p.29-30)

Vocabulary: Period, group, atomic radius, ionization energy.

Equations & Constants: N/A

Practice Problems:

- D 1. The energy required to remove an electron from an atom is the atom's  
 a. electron affinity.  
 b. electron energy.  
 c. electronegativity.  
 d. ☒ ionization energy.

- D 2. The element that has the greatest electronegativity is  
 a. oxygen.  
 b. sodium.  
 c. chlorine.  
 d. ☒ fluorine.

- A 3. One-half the distance between the nuclei of identical atoms that are bonded together is called the  
 a. ☒ atomic radius.  
 b. atomic diameter.  
 c. atomic volume.  
 d. electron cloud.

- C 4. A positive ion is known as a(n)  
 a. ionic radius.  
 b. valence electron.  
 c. ☒ cation.  
 d. anion.

- D 5. A negative ion is known as a(n)  
 a. ionic radius.  
 b. valence electron.  
 c. cation.  
 d. ☒ anion.

- A 6. Across a period in the periodic table, atomic radii  
 a. ☒ gradually decrease.  
 b. gradually decrease, then sharply increase.  
 c. gradually increase.  
 d. gradually increase, then sharply decrease.

- C 7. Which of the following elements has the largest atomic radius?  
 a. Fluorine  
 b. Bromine  
 c. ☒ Iodine  
 d. Chlorine

- A 8. Which of the following elements has the largest atomic radius?  
 a. ☒ Chromium  
 b. Manganese  
 c. Zinc  
 d. Iron

- C 9. Which of the following elements has the largest atomic radius?  
 a. Calcium  
 b. Magnesium  
 c. ☒ Barium  
 d. Strontium

- D 10. Which of the following elements has the largest ionization energy?  
 a. Hydrogen  
 b. Helium  
 c. Lithium  
 d. ☒ Beryllium

- A 11. Which of the following elements has the largest ionization energy?  
 a. ☒ Phosphorous  
 b. Arsenic  
 c. Selenium  
 d. Carbon

- B 12. Which of the following elements has the largest ionization energy?  
 a. Calcium  
 b. ☒ Magnesium  
 c. Barium  
 d. Strontium

#### Atomic Radius

1. Atoms get **larger** going down a group because adding new shells/energy levels

2. Atoms get **smaller** going across a period because not adding energy levels but  $\uparrow Z_{eff}$

# Market Review Packet Answers Key

**Siddappa N.Byrareddy**



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