

MARKING SCHEME

TABLE I

Experiment number	1	ш	III
Final burette reading (cm3)	13.0	25.9	38.7
Initial burette reading (cm3)	0.0	13.0	25.9
Volume of solution B used (cm3)	13.0	12.9	12.8

Complete table	1
Decimal place	1
Accuracy (tied to school value	1
Arithmetic (correct)	1

- (a) Average volume of solution B used 13.0 + 12.9 + 12.8 √½ = 12.9cm³ √½ 3
- (b) Molarity of solution A Molarity = mass / litre = $4.0 = 0.1 \text{M} \sqrt{1}$ RMM 40
- (c) Molarity of solution B $HCl_{(aq)} + NaOH_{(aq)}$ NaCl $_{(aq)} + H_2O_{(d)}$ Mole ratio 1:1 Moles of NaOH = 0.1 x $^{25}/_{1000}$ $\sqrt{1/2}$ =0.0025 moles $\sqrt{1/2}$ Moles of HCl = moles of NaOH = 0.0025 moles Molarity of HCl = 0.0025 x 1,000 $\sqrt{1/2}$ 12.9

TABLE 2

Experiment number		II	LUI
Final burette reading (cm ³)	5.0	10.0	15.0
Initial burette reading (cm3)	0.0	5.0	10.0
Volume of solution B used (cm3)	5.0	5.0	5.0

= 0.19379M √ ½



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RM Cervero

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