



Periodic Table of the Elements

Legend:

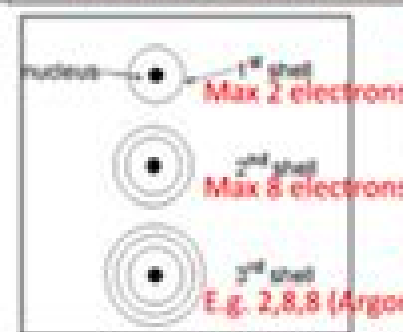
- metals
- non-metals
- metalloids
- transition metals
- alkali metals
- alkali earth metals
- transition metals
- lanthanoids
- actinoids

An atom:
Number of Protons = Number of Electrons
Mass number = Protons + Neutrons

C2 Chemistry

Periodic table arranges elements by
ATOMIC NUMBER (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
E.g. Lithium 2,1 is a group 1 element.

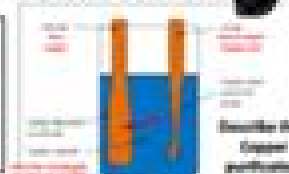


Protons and neutrons both have a mass of 1

ISOTOPES
are atoms of the same element with different numbers of neutrons

OIL RIG

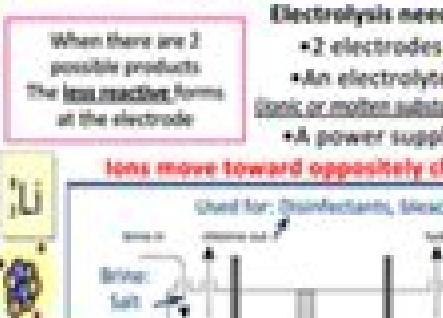
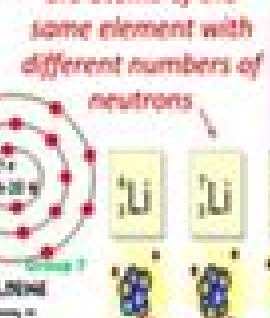
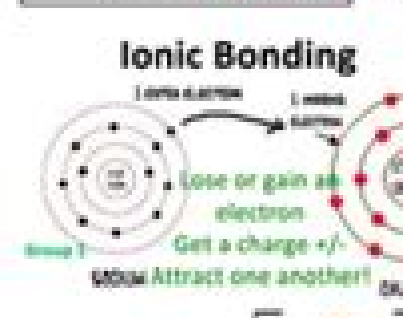
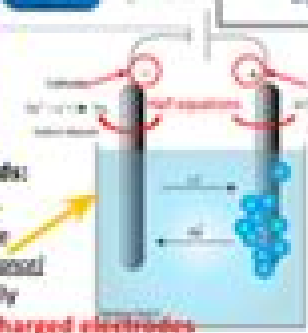
Oxidation is loss (of electrons)
Reduction is gain (of electrons)



Electrolysis:
Splitting up a substance using electricity

Electrolysis needs:

- 2 electrodes
- An electrolyte (ionic or molten substance)
- A power supply



Giant Lattice Structure
Always metals with non-metals
+ ION and - ION attracted by **STRONG ELECTROSTATIC FORCES**

When there are 2 possible products
The less reactive forms at the electrode

As a solid: High melting and Boiling Point
Conducts electricity when melted or dissolved:
Ions move freely and carry a charge

Electrolysis of Brine

Used for: Disinfectants, Bleach and plastics

Used for: Margarine and HCl

Used for: Bleach, paper and soap

NANO-TECHNOLOGY
Means REALLY REALLY REALLY small!!!!

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Table of Contents **Gcse Edexcel Chemistry C2 Past Papers**

1. Understanding the eBook Gcse Edexcel Chemistry C2 Past Papers
 - The Rise of Digital Reading Gcse Edexcel Chemistry C2 Past Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Edexcel Chemistry C2 Past Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gcse Edexcel Chemistry C2 Past Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Edexcel Chemistry C2 Past Papers
 - Personalized Recommendations
 - Gcse Edexcel Chemistry C2 Past Papers User Reviews and Ratings
 - Gcse Edexcel Chemistry C2 Past Papers and Bestseller Lists

5. Accessing Gcse Edexcel Chemistry C2 Past Papers Free and Paid eBooks
 - Gcse Edexcel Chemistry C2 Past Papers Public Domain eBooks
 - Gcse Edexcel Chemistry C2 Past Papers eBook Subscription Services
 - Gcse Edexcel Chemistry C2 Past Papers Budget-Friendly Options
6. Navigating Gcse Edexcel Chemistry C2 Past Papers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gcse Edexcel Chemistry C2 Past Papers Compatibility with Devices
 - Gcse Edexcel Chemistry C2 Past Papers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gcse Edexcel Chemistry C2 Past Papers
 - Highlighting and Note-Taking Gcse Edexcel Chemistry C2 Past Papers
 - Interactive Elements Gcse Edexcel Chemistry C2 Past Papers
8. Staying Engaged with Gcse Edexcel Chemistry C2 Past Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gcse Edexcel Chemistry C2 Past Papers
9. Balancing eBooks and Physical Books Gcse Edexcel Chemistry C2 Past Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gcse Edexcel Chemistry C2 Past Papers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gcse Edexcel Chemistry C2 Past Papers
 - Setting Reading Goals Gcse Edexcel Chemistry C2 Past Papers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gcse Edexcel Chemistry C2 Past Papers
 - Fact-Checking eBook Content of Gcse Edexcel Chemistry C2 Past Papers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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