

Chemistry Honors Worksheet: Formulas of a Compound

- 1.
- moles The subscripts in a chemical formula represent the ratio of (moles / mass) in a compound
 - CO Which compound has the greatest percentage of carbon? (CO, CO₂, PbCO₃)
 - CH₄ What is the empirical formula of a compound that contains 0.2 moles of C and 0.8 moles of H
 - B₃g/mol Calculate the molecular mass of a compound if 0.02461 moles has a mass of 2.05 g.
 - Same The % composition of an empirical formula is (more / less / the same) than/as its molecular formula.
 - one How many moles of V are indicated in one mole of the compound V(C₂O₄)₂?
 - 1×10^{24} atoms A beaker of water contains 2×10^{24} atoms of H, how many atoms of O are also in the beaker?
 - 6.02×10^{23} 1 mole of a molecular compound has a mass of 165 grams. How many molecules are there?
 - 0% Fool's Gold is iron pyrite, FeS. What is the percentage of gold in a 1.0 gram sample of fool's gold?

2. The pesticide DDT has been banned in the United States since 1962. This compound has a formula C₁₄H₉Cl₅. Answer all questions referring to DDT. Note: Use 35.5 g/mol for Cl 709.5 g/mol

- molecular Is this compound ionic or molecular?
- 5.2×10^{15} How many molecules of DDT are in a 6.1×10^{-4} g sample?
- 1.53×10^{17} atoms H How many atoms of H are in 2.00 mg of DDT?
- 7.513 Cl How many grams of Cl are contained in 10.0 grams of DDT?
- 7.84×10^{21} atoms C If a sample 5.04×10^{21} atoms of H, how many atoms of C would there also be?
- 2.97×10^{20} atoms C How many atoms of C are in a 25.0 mg of DDT?

3. A compound of tantalum, Ta, and oxygen was produced in a lab by heating tantalum in a crucible. The data was collected.

Mass of crucible:	24.22 g		
Mass of crucible and tantalum:	26.14 g	<u>> 1.92g Ta</u>	<u>0.0106 mol Ta</u>
Mass of crucible and tantalum oxide:	26.96 g	<u>> 0.42g O</u>	<u>0.02625 mol O</u>

- Calculate the percent composition of the compound. 82.1% Ta 17.9% O
- What is the empirical formula of this compound? TaO_{2.5} - Ta₂O₅
- Is this compound ionic or covalent? ionic
- Name this compound: Tantalum(V) oxide * Need Roman Numerals!!!

4. Determine the empirical formula of the following compounds:

- 63.0g Rb, 5.90g O Rb₂O
- 4.95mg Th, 1.37mg S ThS₂
- 2.13g Na, 2.32g As, 1.98g O Na₃AsO₄
- 32.8% Cr, 67.2% Cl CrCl₃
- 58.0% Rb, 9.50% N, 32.5% O RbNO₃

Empirical Molecular Formula Answers

SA Dillow



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Table of Contents Empirical Molecular Formula Answers

1. Understanding the eBook Empirical Molecular Formula Answers
 - The Rise of Digital Reading Empirical Molecular Formula Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical Molecular Formula Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical Molecular Formula Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical Molecular Formula Answers
 - Personalized Recommendations
 - Empirical Molecular Formula Answers User Reviews and Ratings
 - Empirical Molecular Formula Answers and Bestseller Lists
5. Accessing Empirical Molecular Formula Answers Free and Paid eBooks

- Empirical Molecular Formula Answers Public Domain eBooks
- Empirical Molecular Formula Answers eBook Subscription Services
- Empirical Molecular Formula Answers Budget-Friendly Options
- 6. Navigating Empirical Molecular Formula Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Empirical Molecular Formula Answers Compatibility with Devices
 - Empirical Molecular Formula Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical Molecular Formula Answers
 - Highlighting and Note-Taking Empirical Molecular Formula Answers
 - Interactive Elements Empirical Molecular Formula Answers
- 8. Staying Engaged with Empirical Molecular Formula Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical Molecular Formula Answers
- 9. Balancing eBooks and Physical Books Empirical Molecular Formula Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical Molecular Formula Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical Molecular Formula Answers
 - Setting Reading Goals Empirical Molecular Formula Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical Molecular Formula Answers
 - Fact-Checking eBook Content of Empirical Molecular Formula Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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