

Empirical Formula and Molecular Formula

Lab 6: Molecular Formulas

Lab 6

Students will determine the empirical formula of a compound by determining the mass percent composition of compounds. They will then use the empirical formula and molar mass to determine the molecular formula. This activity will compare an empirical formula to a molecular formula.

Model 1: Comparison of Percent Composition and Empirical Formula

Compound	Empirical Formula	Molecular Formula	Empirical Formula	Percent Composition (Carbon)	Percent Composition (Hydrogen)
1	ethane	C_2H_6	CH_3		
2	propane	C_3H_8	CH_2		
3	ethylene dichloride	$C_2H_2Cl_2$	$CHCl$		
4	benzene	C_6H_6	CH		

1. Looking at the table, what patterns do you see with regard to the percent composition of the compounds in Model 1. What other information do you need to determine the molecular formula?
2. Identify the compounds in Model 1 that have the same empirical formula.
3. Identify the compounds in Model 1 that have the same percent composition.
4. Using the compounds in Model 1, suggest an explanation as to why compounds with different molecular formulas but the same empirical formula have the same percent composition.
5. For each substance in Model 1, give evidence that the molecular formula and empirical formula are exactly twice as far apart as the empirical formula and molecular formula. Give evidence that the molecular formula is exactly twice as far from the empirical formula as the empirical formula.

Empirical Formula And Molecular Formula Pogil Answers

Research and Education Association



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Solving Empirical Formula Problems, 2016 Continue your consideration of classic chemistry problems with a look at empirical formulas and examine how empirical formulas relate to molecular formulas **Minimum Molecular Weight Approach for Determination Empirical Formulas** H.James Harwood, 1965 *Quick Access* Research and Education Association, 2009-09-25 REA's Quick Access Study Charts contain all the information students teachers and professionals need in one handy reference They provide quick easy access to important facts The charts contain commonly used math formulas historical facts language conjugations vocabulary and more Great for exams classroom reference or a quick refresher on the subject **Formulas, Moles and Reactions** Alessio Mangoni, 2025-05-28 This comprehensive and rigorous educational text is dedicated to the study and application of the fundamentals of quantitative chemistry Through sixty clearly and precisely solved exercises it addresses essential topics such as molar mass and molecular weight the concept of the mole Avogadro's number the calculation of atoms and molecules in known quantities percent composition of compounds determination of empirical and molecular formulas and the balancing of chemical equations It serves as an effective tool for chemistry students and for anyone wishing to deepen their understanding of the discipline through a practical and methodical approach providing solid skills that are essential for continuing in scientific studies [Empirical Formula and Ligand Formula Index for Parts 1 to 4](#) Leopold Gmelin, Richard Josef Meyer, Gmelin Institut für Anorganische Chemie und Grenzgebiete in der Max-Planck-Gesellschaft zur Förderung der Wissenschaften, Deutsche Chemische Gesellschaft, 1987 **Chemical Formulas and Names** Virginia P. Powell, 1965 [Empirical Formulas](#) Theodore Rudolph Running, 2006 **Empirical Formulas** Theodore Rudolph Running, 1917 [Empirical Formulas](#) Th.R. Running, 1917

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