

EMPIRICAL AND MOLECULAR FORMULA

- 1) What is the empirical formula of a compound that contains 46.3 % lithium and 53.7% oxygen?
- 2) What is the empirical formula of a compound that contains 15.9 % boron and 84.1 % fluorine?
- 3) Phosphorus reacts with oxygen to give a compound that is 43.7% phosphorus and 56.4% oxygen. What is the empirical formula of the compound?
- 4) An inorganic salt is composed of 17.6% sodium, 39.7% chromium and 42.8% oxygen. What is the empirical formula of this salt?
- 5) Compound X contains 69.9% carbon, 6.86% hydrogen and 23.3% oxygen. Determine the empirical formula of X.
- 6) Oxalic acid has the empirical formula CHO_2 . Its molar mass is 90 g/mol. What is the molecular formula of oxalic acid?
- 7) The empirical formula of codeine is $\text{C}_{18}\text{H}_{21}\text{NO}_3$. If the molar mass of codeine is 299 g/mol, what is its molecular formula?
- 8) A compound's molar mass is 240.28 g/mol. Its percentage composition is 75.0% carbon, 5.05% hydrogen and 20.0% oxygen. What is the compound's molecular formula?
- 9) The wintergreen plant produces methyl salicylate, or oil of wintergreen. It can also be prepared easily in the laboratory. Methyl salicylate is 63.1% carbon, 5.31% hydrogen and 31.6% oxygen. Calculate the empirical formula of this compound.
- 10) An inorganic salt is made up of 38.8% calcium, 20.0% phosphorus, and 41.2% oxygen.
 - a) What is the empirical formula of the compound?
 - b) On further analysis, each formula unit of this salt is found to contain two phosphate ions. Predict the molecular formula of this salt.

Answers

- 1) Li_2O
- 2) BF_3
- 3) P_2O_5
- 4) $\text{Na}_2\text{Cr}_2\text{O}_7$
- 5) $\text{C}_{12}\text{H}_{14}\text{O}_3$
- 6) $\text{C}_2\text{H}_2\text{O}_4$
- 7) $\text{C}_{18}\text{H}_{21}\text{NO}_3$
- 8) $\text{C}_{15}\text{H}_{12}\text{O}_3$
- 9) $\text{C}_8\text{H}_8\text{O}_3$
- 10) a) $\text{Ca}_3\text{P}_2\text{O}_8$ b) $\text{Ca}_3(\text{PO}_4)_2$

Empirical And Molecular Formula Problems

Sabine Zange

A decorative red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

Empirical And Molecular Formula Problems:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Solving Empirical Formula Problems*, 2016 Continue your consideration of classic chemistry problems with a look at empirical formulas and examine how empirical formulas relate to molecular formulas Chemistry: 1,001 Practice Problems For Dummies (+ Free Online Practice) Heather Hattori, Richard H. Langley, 2014-03-11 Practice makes perfect and helps deepen your understanding of chemistry Every high school requires a course in chemistry and many universities require the course for majors in medicine engineering biology and various other sciences 1001 Chemistry Practice Problems For Dummies provides students of this popular course the chance to practice what they learn in class deepening their understanding of the material and allowing for supplemental explanation of difficult topics 1001 Chemistry Practice Problems For Dummies takes you beyond the instruction and guidance offered in Chemistry For Dummies giving you 1 001 opportunities to practice solving problems from the major topics in chemistry Plus an online component provides you with a collection of chemistry problems presented in multiple choice format to further help you test your skills as you go Gives you a chance to practice and reinforce the skills you learn in chemistry class Helps you refine your understanding of chemistry Practice problems with answer explanations that detail every step of every problem Whether you re studying chemistry at the high school college or graduate level the practice problems in 1001 Chemistry Practice Problems For Dummies range in areas of difficulty and style providing you with the practice help you need to score high at exam time

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High School Chemistry Michael A. Buben, 2018-05-18 High School Chemistry is an outline of notes It was provided to chemistry students so they could focus on the material being presented and not be burdened by drawing charts copying definitions writing problems etc Mr Buben added his instructional notes via a Teacher s Edition The books go hand in hand Modified Second Edition as a student s workbook and a Teacher s Edition for guidance

Comprehensive Chemistry XI Dr. B. Kapila, S. K. Khanna, 2010-11 Comprehensive chemistry according to the new syllabus prescribed by Central Board of Secondary Education CBSE

SELF-HELP TO SIMPLIFIED CHEMISTRY 10 (FOR 2022-23 EXAMINATIONS) Amar Bhutani, This book is written strictly in accordance with the latest syllabus prescribed by the Council for the I C S E Examinations in and after 2023 This book includes the Answers to the Questions given in the Textbook

Simplified Chemistry Class 10 published by Allied Publications Pvt Ltd This book is written by Amar Bhutani

This book delves into Empirical And Molecular Formula Problems. Empirical And Molecular Formula Problems is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Empirical And Molecular Formula Problems, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Empirical And Molecular Formula Problems
- Chapter 2: Essential Elements of Empirical And Molecular Formula Problems
- Chapter 3: Empirical And Molecular Formula Problems in Everyday Life
- Chapter 4: Empirical And Molecular Formula Problems in Specific Contexts
- Chapter 5: Conclusion

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3. In chapter 2, this book will delve into the foundational concepts of Empirical And Molecular Formula Problems. The second chapter will elucidate the essential principles that must be understood to grasp Empirical And Molecular Formula Problems in its entirety.

4. In chapter 3, the author will examine the practical applications of Empirical And Molecular Formula Problems in daily life. This chapter will showcase real-world examples of how Empirical And Molecular Formula Problems can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Empirical And Molecular Formula Problems in specific contexts. The fourth chapter will explore how Empirical And Molecular Formula Problems is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Empirical And Molecular Formula Problems. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Empirical And Molecular Formula Problems.

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