

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of C and 6.67 g of H. Find the empirical formula.

$$\text{C} = 40.0 \text{ g} \times \frac{1 \text{ mol}}{12.01 \text{ g}} = 3.33 \text{ mol}$$

$$\text{H} = 6.67 \text{ g} \times \frac{1 \text{ mol}}{1.01 \text{ g}} = 6.60 \text{ mol}$$

C₃H₆

2. A compound is found to contain 50.0 g of Fe, 22.0 g of S, and 28.0 g of O. Find the empirical formula.

$$\text{Fe} = 50.0 \text{ g} \times \frac{1 \text{ mol}}{55.85 \text{ g}} = 0.895 \text{ mol}$$

$$\text{S} = 22.0 \text{ g} \times \frac{1 \text{ mol}}{32.07 \text{ g}} = 0.686 \text{ mol}$$

$$\text{O} = 28.0 \text{ g} \times \frac{1 \text{ mol}}{16.00 \text{ g}} = 1.75 \text{ mol}$$



3. A compound is found to contain 40.0 g of C, 6.67 g of H, and 33.33 g of O. Find the empirical formula.

$$\text{C} = 40.0 \text{ g} \times \frac{1 \text{ mol}}{12.01 \text{ g}} = 3.33 \text{ mol}$$

$$\text{H} = 6.67 \text{ g} \times \frac{1 \text{ mol}}{1.01 \text{ g}} = 6.60 \text{ mol}$$

$$\text{O} = 33.33 \text{ g} \times \frac{1 \text{ mol}}{16.00 \text{ g}} = 2.08 \text{ mol}$$



4. A compound is found to contain 40.0 g of C, 6.67 g of H, and 53.33 g of O. Find the empirical formula.

$$\text{C} = 40.0 \text{ g} \times \frac{1 \text{ mol}}{12.01 \text{ g}} = 3.33 \text{ mol}$$

$$\text{H} = 6.67 \text{ g} \times \frac{1 \text{ mol}}{1.01 \text{ g}} = 6.60 \text{ mol}$$

$$\text{O} = 53.33 \text{ g} \times \frac{1 \text{ mol}}{16.00 \text{ g}} = 3.33 \text{ mol}$$



Empirical Formula Pogil Science Division Bellevue

Annelies Wilder-Smith



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Decoding **Empirical Formula Pogil Science Division Bellevue**: Revealing the Captivating Potential of Verbal Expression

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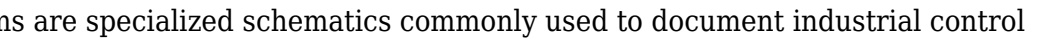


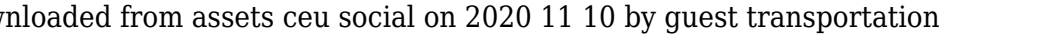


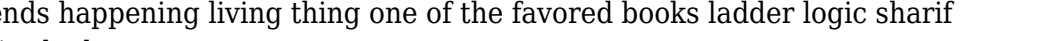

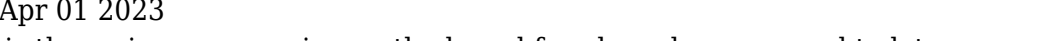
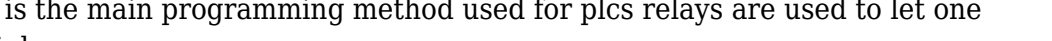



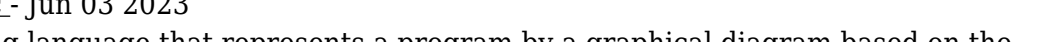
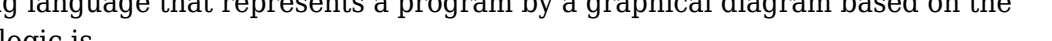

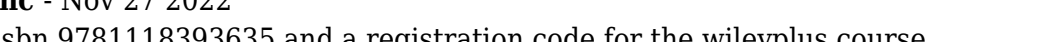
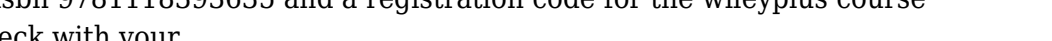

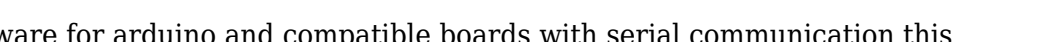
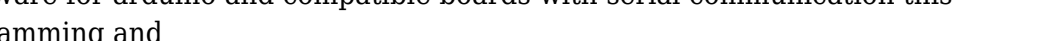




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