

# Empirical and Molecular Formulas Worksheet

Key

## Objectives:

- be able to calculate empirical and molecular formulas

## Empirical Formula

- 1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l} \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} = 1.96 \quad 2 \\ \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} = 5.88 \quad 6 \\ \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1 \end{array}$$



- 2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5}$$

$$\text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5$$



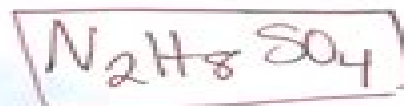
- 3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad 2$$

$$\text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad 8$$

$$\text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad 1$$

$$\text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad 4$$



## Molecular Formula

- 4) Empirical formula of a substance is  $\text{CH}_2\text{O}$ . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 1 + 16) =$$

$$\frac{180}{30} = 6$$



- 5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

$$\text{C} \quad \frac{1.388}{3.585} \quad 38.7\% \rightarrow \frac{3.88\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{H} \quad \frac{0.345}{3.585} \quad \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3$$

$$\text{O} \quad \frac{1.850}{3.585} \quad \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 1 + 16)$$

$$\frac{62}{31.03} = 2$$



1	$\text{C}_2\text{H}_6\text{O}$
2	$\text{Au}_2\text{O}_3$
3	$\text{NH}_4\text{SO}_4$
4	$\text{C}_2\text{H}_6\text{O}_2$
5	$\text{C}_6\text{H}_{12}\text{O}_6$

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## **Table of Contents Empirical And Molecular Formula Practice Answer Key**

1. Understanding the eBook Empirical And Molecular Formula Practice Answer Key
  - The Rise of Digital Reading Empirical And Molecular Formula Practice Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Empirical And Molecular Formula Practice Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Empirical And Molecular Formula Practice Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical And Molecular Formula Practice Answer Key
  - Personalized Recommendations
  - Empirical And Molecular Formula Practice Answer Key User Reviews and Ratings
  - Empirical And Molecular Formula Practice Answer Key and Bestseller Lists
5. Accessing Empirical And Molecular Formula Practice Answer Key Free and Paid eBooks

- Empirical And Molecular Formula Practice Answer Key Public Domain eBooks
- Empirical And Molecular Formula Practice Answer Key eBook Subscription Services
- Empirical And Molecular Formula Practice Answer Key Budget-Friendly Options
- 6. Navigating Empirical And Molecular Formula Practice Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Empirical And Molecular Formula Practice Answer Key Compatibility with Devices
  - Empirical And Molecular Formula Practice Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Empirical And Molecular Formula Practice Answer Key
  - Highlighting and Note-Taking Empirical And Molecular Formula Practice Answer Key
  - Interactive Elements Empirical And Molecular Formula Practice Answer Key
- 8. Staying Engaged with Empirical And Molecular Formula Practice Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Empirical And Molecular Formula Practice Answer Key
- 9. Balancing eBooks and Physical Books Empirical And Molecular Formula Practice Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Empirical And Molecular Formula Practice Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical And Molecular Formula Practice Answer Key
  - Setting Reading Goals Empirical And Molecular Formula Practice Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical And Molecular Formula Practice Answer Key
  - Fact-Checking eBook Content of Empirical And Molecular Formula Practice Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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