

Pre-lab Questions: The Determination of the Empirical Formula of a Hydrate

A **hydrate** is an ionic compound (ionic salt) that contains water molecules in its structure. An **anhydride** (also called anhydrous salt) is the substance that remains after the water is removed from a hydrate. When a hydrate is heated, the water molecules are removed in the form of steam, leaving behind the anhydrite. To determine the formula of a hydrate experimentally, the **mole ratio** of the water portion to the mole ratio of the anhydrous portion must be calculated.

Questions:

During a Chem 121 experiment, a student obtained the following data after heating a sample of calcium chloride hydrate $\text{CaCl}_2 \cdot x\text{H}_2\text{O}$, where x refers to the number of water molecules hydrated per molecule of CaCl_2 .

Data:

Weight of crucible and lid (g)	32.1070
Weight of crucible, lid and hydrated salt (g)	36.8283
Weight of crucible, lid and anhydrous salt after	
a) 1 st heating (g)	35.7347
b) 2 nd heating (g)	35.6682
c) 3 rd heating (g)	35.6679

Be sure to show the general formula used for each step.

1. Use this information to determine the weight of the hydrated salt ($\text{CaCl}_2 \cdot x\text{H}_2\text{O}$):

- a) the weight of the anhydrous salt (CaCl_2)

Note: the weight after the 2nd and 3rd heating are close enough in value to assume that constant weight has been reached and all the water molecules have been removed

- b) the weight of water removed by heating

Empirical Formula Determination Pre Lab Question Answers

Emilie Sanchez

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