## Question 1

- 1(b)(i) Reaction rate decreased as time increased.
- 1(b)(ii) The smaller zinc particles gave a larger surface area which led to a faster reaction rate since there were more successful collisions per unit time.
- 1(b)(iii) Temperature, pressure (for gases) and concentration (of reagents in solution).
- 1(c) One safety precaution could have been the wearing of safety glasses/goggles to protect the eyes from possible acid spray.
- 1(d) Temperature could have been controlled during the experiment by the use of a water bath.
- 1(e)(i)  $Zn_{(a)} + H_2SO_{a(aa)} \rightarrow ZnSO_{4(aa)} + H_{J(a)}$
- 1(e)(ii) From the balanced equation for the reaction. 1 mole of Zn<sub>(b)</sub> produces 1 mole of H<sub>2(g)</sub>. Moles of zinc used = mass of Zn used/molar mass of Zn = (0.55/65.4) mol. i.e. 0.0084 mol. Zn. Therefore 0.0084 moles of H<sub>2(g)</sub> were produced. Volume of H<sub>2(g)</sub> produced = moles of H<sub>2(g)</sub> produced × molar volume of gas at RTP. Volume of H<sub>2(g)</sub> produced = (0.0084 × 24 000) cm<sup>3</sup> = 201.8 cm<sup>3</sup>.

## Question 2

- 2(a) The number of occupied electron shells in an tells us what period it is in. All these elements have three occupied electron shells and are therefore in period three.
- 2(b) Element: Sodium (Na).

Equation:  $Na_2O_{(i)} + H_2O_{(i)} \rightarrow 2NaOH_{(aq)}$ .

2(c) Allotrope – one of multiple (2 or more) different forms/structures of the same element, which exist in the same physical state.



- 2(e)(i) When sulfur reacts with chlorine there is covalent bonding.
- 2(e)(ii) Both sulfur and chlorine are non-metals. It is energetically preferable for these atoms to share electrons with each other in order to achieve stable outer electronic configurations of 8 electrons, rather than to lose or gain electrons.
- 2(f) SCl<sub>2</sub>.
- S sulfur nucleus and inner electrons. × sulfur electron.

  (CI) chlorine nucleus and inner electrons. – chlorine electron.

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**JG Myers** 

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