

Question Number	Acceptable Answers	Reject	Mark
16(b)	First mark $\text{Moles NaOH} = \frac{0.00750}{1000} \times 20.0$ $= 1.50 \times 10^{-4} \text{ (mol)}$ (1) (Since HCOOH : NaOH ratio is 1:1) Second mark $[\text{HCOOH(aq)}] = \frac{1.50 \times 10^{-4}}{0.0250}$ OR $= 1.50 \times 10^{-4} \times \frac{1000}{25.0}$ (1) $(\text{= } 6.00 \times 10^{-3} \text{ mol dm}^{-3})$ ALTERNATIVE APPROACH: Use of an expression such as $0.00750 \times 20.0 = 25 \times y$ (1) $y = \frac{0.00750 \times 20.0}{25}$ (1)		2

Question Number	Acceptable Answers	Reject	Mark
16(c)(i)	(Weak) dissociates / ionizes to a small extent OR dissociate / ionizes partially OR dissociates / ionizes incompletely OR does not fully dissociate / ionize OR forms an equilibrium when reacted with water (1) (Acid) proton donor ALLOW 'proton donator' OR produces / releases H^+ ions OR produces / releases H_3O^+ ions (1) Ignore reference to typical acid reactions	'not easily dissociated'	2

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RJ Shavelson



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