

Half-Modern Chemistry Review

CHAPTER 11: CHEMICAL BONDING

This half-modern program concludes that level (that vast 50%) of that information for this chapter is dead.

Focus on this content, but make sure to review class notes, activities, materials, experiments, etc.

If you study this document and DISCOVERED what, you should at least be able to PLACE that item.

Now that items will be recall, examples, and/or application of this content. Now

OUTCOMES

- Communicate with precision the understanding/properties learned (p2 c2)
- Communicate understanding learned to teachers and/or peers (p2 c2)
- EL-2 Give the Lewis dot notation, illustrate the placement bond on the drawing of molecules (p2 d, p2)
- EL-2 List important physical properties of some compounds; named molecules/molecules in some (p2 d, p2)

EL-1: INTRODUCTION TO CHEMICAL BONDING

• Introduction

- a) chemical bond → the attraction force that binds atoms or ions together
- a) ionic bonding → a force that attracts electrons from one atom to another, which demonstrates a mutual attraction between them
- a) covalent bonding → a bond formed when atoms share electrons with one another, at least one.
- a) intermolecular forces → a bond formed by attractive forces between atoms and molecules, which are the basic chemical bonds.
- a) gas → substance, a molecule or atoms that particles are separated by large and irregular spaces and compressible
- a) gas-liquid bond → a chemical bond in which a gas and liquid are present together, at some or both ends of a system, they are also

• Chapter Highlights

- a) ionic bonds are characteristically formed by ionic atoms. They show major types of chemical bonding, and have, usually, and usually.
- a) In general, atoms of metals, bond covalently with atoms of nonmetals, atoms, or metals bond covalently with each other, and atoms of nonmetals bond covalently with each other.

EL-2: COVALENT BONDING AND MOLECULAR COMPOUNDS

• Introduction

- a) molecule → a group of atoms that are held together by chemical bonds, a molecule is the smallest unit of matter that can exist by itself and retain all of its chemical properties
- a) molecular compound → a chemical compound whose simplest unit, and molecules
- a) chemical formula → a combination of chemical symbols and numbers to represent a substance
- a) molecular formula → a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms

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