CHAPTER - 00

REDOX REACTIONS AND ELECTROCHEMISTRY

EARLIER CONCEPT OF OXIDATION

Oxidation is the process of addition of oxygen or any electronegative element or group or removal of hydrogen or any electropositive element of group.

Eg :
$$2Mg + O_2 \longrightarrow 2MgO$$

 $2FeCl_2 + Cl_2 \longrightarrow 2FeCl_3$
 $H_2S + Cl_2 \longrightarrow 2HCl + S$
 $K_3MnO_4 + O_3 + H_3O \longrightarrow KMnO_4 + KOH$

Reduction is the process of addition of hydrogen or any electropositive element or group or removal of oxygen or any electronegative element or group.

Eg:
$$H_2 + Cl_2 \longrightarrow 2HCl$$

 $2HgCl_2 + SnCl_2 \longrightarrow Hg_3Cl_2 + SnCl_4$
 $ZnO + C \longrightarrow Zn + CO$
 $2FeCl_1 + H_2S \longrightarrow FeCl_2 + 2HCl + S$

ELECTRONIC CONCEPT

According to the electronic concept, oxidation involves loss of electrons or removal of electrons.

Eg :
$$MnO_4^2 \longrightarrow MnO_4^- + 1e^-$$

 $Fe^{2+} \longrightarrow Fe^{3+} + 1e^-$
 $Mg \longrightarrow Mg^{2+} + 2e^-$

Here MnO_{a}^{2} , Fe^{2} and Mg are donating electrons (electron donor) and are called Reducing agents.

... Reducing agent is electron donor or we can say that reducing agents are getting oxidised. Similarly

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