

Fischer Esterification Project

Introduction

Esters happens to be chemical compounds that play a very significant role in the lives of people worldwide. According to *Encyclopedia.com*, naturally esters of organic acids in fruits and flowers enables them to have their distinctive odors. Mostly, they are synthesized in the laboratory and used by cosmetic industries. Nonetheless, certain evaporative esters are used as solvents for paints, lacquers and varnishes. Scientifically, common fats and oils are combination of various esters like stearin, palmitin, and linolein, formed from the alcohol glycerol and fatty acids. (*Encyclopedia.com*)

On the other hand, Fischer Esterification is a method used in synthesizing esters from carboxylic acids and alcohols. As depicted in Fig 1, it occurs with the liberation of water. The reaction is acid catalyzed and an inorganic acid as sulfuric acid can be used as a catalyst. [Williamson, 2011]. According to Le Chatelier's Principle, the reaction is reversible since the position of equilibrium can be shifted to the right by either the addition or removal of excess reactants.



Fig 1: General representation of Fischer Esterification [McMurry, 2016]

The mechanism as shown in (Fig 2) happens by the initial protonation of carboxyl group by the acid, followed by nucleophilic attack by the alcohol, proton rearrangement, loss of water and loss of catalyzing proton. The reversible reaction yields hydrolysis of the ester, [Williamson, 2011]



Fig 2: General Mechanism for Fischer Esterification [Williamson, 2011]

Objectives

For this experiment, unknown samples of a carboxylic acid and an alcohol were given for the Fischer ester synthesis. Prior to the experiment, the molar masses of the reactants to be used were given and their refractive indices were taken so all the substances were known.

Procedure

Esterification Experiment Report

**Connecticut Agricultural Experiment
Station**



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