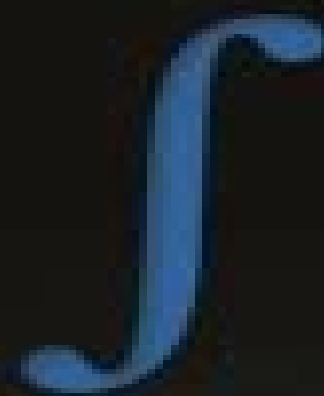




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Essential Calculus Early Transcendentals

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subsequent ΔH values $2 \text{ CO}_2(g) + \text{H}_2\text{O}(g) \rightarrow \text{C}_2\text{H}_2(g) + \frac{5}{2} \text{O}_2(g)$ $\Delta H = 94.5 \text{ kJ}$ $\text{H}_2\text{O}(g) \rightarrow \text{H}_2(g) + \frac{1}{2} \text{O}_2(g)$ $\Delta H = 71.2 \text{ kJ}$ $\text{C}_2\text{H}_6(g) \rightarrow \frac{7}{2} \text{O}_2(g) + 2 \text{CO}_2(g) + 3 \text{H}_2\text{O}(g)$ $\Delta H = 283 \text{ kJ}$

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web thus for reaction 9 2 in which 1 mole of h₃o and 1 mole of oh combine to form 2 moles of h₂o h 55 8 kj as shown note that the enthalpy change h h_{products} h_{reactants} is positive if heat is absorbed that is if h_{products}

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