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User's Guide

Mettler Toledo 8461 Manual

Michal Rosen-Zvi

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with 1200 ml of 2M HNO_3 if the final density of solution is same as pure water then molarity of the iron in resulting solution which decides the nature of the above solution is 1 0 25M 2 0 5M 3 0 888M 4 none of

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$$m = \frac{\text{moles of solute}}{\text{kg of solvent}}$$

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web 1 molarity $M = \frac{m}{V}$ mass molar mass $\times 0.4500 \text{ L} = 0.825 \text{ g}$ 141.9579 g/mol $\times 0.0129 \text{ M} = 2 \text{ molality}$ 0.825 g 141.9579 g/mol 0.00581158 mol 0.00581158 mol 0.4500 kg 0.0129 M 3 mole fraction $n_2 = \frac{0.825 \text{ g}}{141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ $n_1 = \frac{450.0 \text{ g}}{18.015 \text{ g/mol}} = 24.97918401 \text{ mol}$

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solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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$$M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_1 + V_2 + \dots + V_n}$$
 with increase in temperature which of these changes a molality b fraction of solute present in water

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