

1 Determine the volume of CO_2 gas given off when excess MgCO_3 is added to 25.0 cm^3 of 0.400 mol/dm^3 HCl at room temperature and pressure.



Use the following steps.

- Calculate the number of moles of HCl in 25.0 cm^3 of 0.400 mol/dm^3 of acid.

$$c = \frac{m}{v}$$

$$m = 0.4 \times 0.025 \\ = 0.01$$

.....0.01..... mol

- Determine the number of moles of CO_2 gas given off.



$$0.001 : x$$

$$0.01 \div 2 = 0.005$$

.....0.005..... mol

- Calculate the volume of CO_2 gas given off in cm^3 .

$$1 \text{ mol} = 24 \text{ dm}^3$$

$$0.005 \text{ mol} = x \text{ dm}^3$$

$$x = 0.005 \times 24$$

$$= 0.12 \text{ dm}^3$$

$$= 120 \text{ cm}^3 (\times 1000)$$

.....120..... cm^3
[3]

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