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 $550 \text{ J} = \Delta h \times q_{\text{water}} = 550 \text{ J}$

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web calorimetry problems calorimetry problems $q_{\text{sur}} = m \times c \times \Delta t$ heat m mass q_{rxn} q_{sur} t_f t_i name per date c specific heat
 for water $4184 \text{ J/g} \times 1$ what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased
 by $8.1 \text{ }^{\circ}\text{C}$ when 207 J of heat is added 2

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 $25.0 \text{ }^{\circ}\text{C}$ and heat it to $145.0 \text{ }^{\circ}\text{C}$ what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the
 process in 1 3 i have 18.7 grams of an unknown liquid chemical at a temperature of $35.0 \text{ }^{\circ}\text{C}$

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 of 1.00 M NaOH at $25.0 \text{ }^{\circ}\text{C}$ in a foam cup calorimeter a reaction occurs and the temperature rises to $32.0 \text{ }^{\circ}\text{C}$ calculate the
 enthalpy change in kJ for this reaction assume the densities of the solutions are 1.00 g/ml and that

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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web 6 a 218 g sample of steam at 121 $^\circ\text{C}$ is cooled to ice at 14 $^\circ\text{C}$ find the change in heat content of the system 7 in going from ice at 34 $^\circ\text{C}$ to steam at 138 $^\circ\text{C}$ a sample of water absorbs $1 41 \times 10^5$ J find the mass of the sample answers 1 2 68 $\times 10^4$ J 2 6 23 $\times 10^4$ J 3 1 71 $\times 10^4$ J 4 1 11 $\times 10^5$ J 5 2 11 $\times 10^5$ J 6

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