

CHEMISTRY



Periodic Table of Elements

1																	18	
H																	Ar	
2																	19	
He																	K	
3	4															20	36	
Li	Be															Ca	Kr	
5	6	7	8	9	10												37	
B	C	N	O	F	Ne												Rb	
11	12															38	54	
Na	Mg															Sr	Xe	
13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	36
Al	Si	P	S	Cl	Ar	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Kr
31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48	54
Ga	Ge	As	Se	Br	Kr	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	Xe
49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	72
In	Sn	Sb	Te	I	Xe	Cs	Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Hf	
69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	94
Tl	Pb	Bi	Po	At	Rn	Fr	Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	
95	96	97	98	99	100	101	102	103	104	105	106	107	108	109	110	111	112	118
Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uue
119	120	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136	142
Uut	Uuq	Uub	Uut	Uuq	Uub	Uut	Uuq	Uub	Uut	Uuq	Uub	Uut	Uuq	Uub	Uut	Uuq	Uub	Uut
143	144	145	146	147	148	149	150	151	152	153	154	155	156	157	158	159	160	166
Uuq	Uub	Uut	Uuq	Uub	Uut	Uuq	Uub	Uut	Uuq	Uub	Uut	Uuq	Uub	Uut	Uuq	Uub	Uut	Uuq

1	ATOMIC NUMBER
H	SYMBOL
1.008	ATOMIC WEIGHT

- Elemental forms:**
 - Monoatomic:** Noble gases: helium, neon, argon, krypton, xenon, and radon; nonreactive
 - Diatoms:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid: bromine; Solid: iodine
 - Molecular solids:** Sulfur
 - Covalent solids:** Carbon (diamond, graphite)
 - Metals:** Most common type of element; high coordination numbers; very dense
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options
 - Ionic:** Electron exchange
 - Covalent:** Shared electrons
 - Polar covalent: Unequal sharing
 - Ion complex:** Central ion + Solvent
 - Transition metal compounds:**
 - Coordination compounds: Non-bonding ligands
 - Organometallics: Metal-C bond
 - Complex minerals:** Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
 - Family of elements:** Column in periodic table
 - The addition of d-orbitals produces an expanded octet

Periodic Trends

- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cation and anions (same charge) decrease in size as the atomic number (Z) increases

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens Z -effect on orbital energy
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme
 - Ionic radius:** Depends on charge, counter ion, and coordination (crystal type)
 - Atomic radius:** Based on diatomic gas data or metallic solid structural data

Development of the Periodic Table

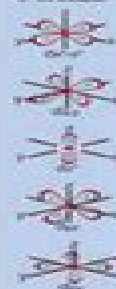
- Missing elements were predicted using periodicity
- Column 1: Alkali metals
- Column 2: Alkaline earth metals
- First transition series: Titanium-Copper
- Column 17: Halogens: fluorine, chlorine, bromine, iodine, and astatine
- Column 18: Rare gases: helium, neon, argon, krypton, xenon, and radon
- Actinides: $Z = 90$ or above; radioactive
- Lanthanides: f-subshell elements; starts with $Z = 58$



Practical Application

- The colors of fireworks depend on the elements in the explosive

d-subshell 5 orbitals



p-subshell 3 orbitals



s-subshell 1 orbital



Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge $-q$, $Z = +q$ electrons; for a cation of charge $+q$, $Z = -q$ electrons
- Mass number (A):** $A = Z + N$, where N = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same Z , different A ; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)



Atomic Quantum Numbers & Orbitals

- n : principle**
 - angular momentum (orbital shape)
 - quantum number (orbital direction)
- l : orbital**
 - quantum number (orbital shape)
 - quantum number (orbital direction)
- m_l : magnetic**
 - quantum number (orbital direction)
 - quantum number (orbital shape)
- m_s : spin**
 - quantum number (orbital direction)
 - quantum number (orbital shape)

Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:

Aufbau Order for Filling Subshells



n	l	m_l	m_s	element
1	0	0	$+\frac{1}{2}$	H
1	0	0	$-\frac{1}{2}$	He
2	0	0	$+\frac{1}{2}$	Li
2	0	0	$-\frac{1}{2}$	Be
2	1	-1	$+\frac{1}{2}$	B
2	1	0	$+\frac{1}{2}$	C
2	1	1	$+\frac{1}{2}$	N
2	1	-1	$-\frac{1}{2}$	O
2	1	0	$-\frac{1}{2}$	F
2	1	1	$-\frac{1}{2}$	Ne

- Electron configuration:** Atomic orbital occupancy for an atom
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons: one with spin up, $m_s = +\frac{1}{2}$, and one with spin down, $m_s = -\frac{1}{2}$
- Hund's rule:** Electrons fill p, d, and f subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals); spin is maximized
- Ionization potential (IP):** Energy required to remove electron from an atom or ion
 - First IP:** Removal of first valence electron
- Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 0 for rare gas

Lavc Chemistry Study Guide

G Psacharopoulos



Lavc Chemistry Study Guide:

This book delves into Lavc Chemistry Study Guide. Lavc Chemistry Study Guide is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Lavc Chemistry Study Guide, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Lavc Chemistry Study Guide
- Chapter 2: Essential Elements of Lavc Chemistry Study Guide
- Chapter 3: Lavc Chemistry Study Guide in Everyday Life
- Chapter 4: Lavc Chemistry Study Guide in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Lavc Chemistry Study Guide. The first chapter will explore what Lavc Chemistry Study Guide is, why Lavc Chemistry Study Guide is vital, and how to effectively learn about Lavc Chemistry Study Guide.
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6. In chapter 5, this book will draw a conclusion about Lavc Chemistry Study Guide. This chapter will summarize the key points that have been discussed throughout the book.

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