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Abstract

- ### ३. मरु-संस्कृत संज्ञा

☞ কোলাস পানিসূত্র লবণের দ্রাব্যতা আপনাত্মা কৃষ্ণির সাথে সাথে এক পর্যায়ে কৃষ্ণি না পেয়ে হ্রাস পায়। এর কারণ, আপনাত্মা কৃষ্ণির সাথে সাথে দ্রাব্যতা কৃষ্ণি পেতে থাকলেও একটি নির্দিষ্ট আপনাত্মার কোলাস পানি দ্রুত লবণ নিষ্কৃতিত হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে দ্রবীভূত হওয়ার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরী বিক্রিয়া কিন্তু 32.4°C আপনাত্মার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C আপনাত্মার উপরে আপনাত্মা কৃষ্ণিজনিত কারণে দ্রাব্যতা হ্রাস ঘটে।

w = SMV
 = $(0.5 \times 98 \times 0.25) \text{ g}$
 = 12.25 g
 অতএব, শুষ্কত H₂SO₄ এর ভর
 = 12.25 g

উদাহরণের A-পাত্রে অবস্থিত,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$

দ্রবণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
 দ্রবণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

સરીકૂટ H_2SO_4 એર ભર, $w = 7$

- নোড কাকে বলে?
- Cl_2 অণুজার— ব্যাখ্যা করে।
- উদ্ভীপকের এসিডটির মোলার ঘনমাত্রা নির্ণয় করে।
- উদ্ভীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করে।

२. नगर कार्यपालिका छिन्दा

১৩. সমযোজী বন্ধনে আনন্দ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সমান হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য ০.৫ এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা ৩.০ ও C অপুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_4 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য মৌলিক অণুর ক্ষেত্রে এছাড়া থাকে।

গ) উদ্দীপকে বর্ণিত এসিডটি হলো $2\% \left(\frac{W}{V}\right) HCl$

2% $\left(\frac{W}{V}\right)$ HCl এর অর্থ হলো—

100 mL. ক্ষরণে HCl এর ভর 2g

$$\therefore 1000 \text{ mL} = \left(\frac{2 \times 1000}{100} \right) \text{ g} = 20 \text{ g}$$

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