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**MEMORANDUM**

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**KJ Lindholm-Leary**



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web jan 29 2020 the ionic form of this reaction is  $2\text{Na}^+\text{aq} + 2\text{Cl}^-\text{aq} + \text{Cu}^{2+}\text{aq} + \text{SO}_4^{2-}\text{aq} \rightarrow 2\text{Na}^+\text{aq} + \text{SO}_4^{2-}\text{aq} + \text{Cu}^{2+}\text{aq} + 2\text{S}^{2-}\text{aq}$  the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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net ionic equation  $\text{H}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{CO}_2$  here is a video on spectator ions  
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web the six  $\text{Na}^+$  ions and the six  $\text{Cl}^-$  ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation  $3\text{Ba}^{2+}\text{aq} + 2\text{PO}_4^{3-}\text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2\text{s}$

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 $\text{AgNO}_3\text{aq} + \text{KClaq} \rightarrow \text{AgCl s} + \text{KNO}_3\text{aq}$  total ionic  $\text{Ag}^+\text{aq} + \text{NO}_3^-\text{aq} + \text{K}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl s} + \text{K}^+\text{aq} + \text{NO}_3^-\text{aq}$  net ionic  $\text{Ag}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl s}$  2

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{H}^+\text{aq} + \text{OH}^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{l}$  the spectator ions in  $\text{KOH aq}$  and  $\text{HNO}_3\text{aq}$  are  $\text{K}^+\text{aq}$  and  $\text{NO}_3^-\text{aq}$  how to identify and list the

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