

Year 11 C2 - Discovering Chemistry Topic 1

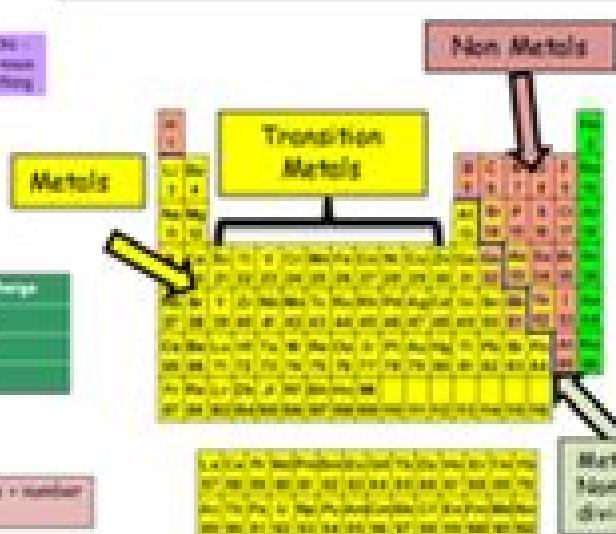


Particle	Relative Mass	Relative Charge
Proton	1	+
Neutron	1	0
Electron	Very small	-

I came up with the first periodic table. I arranged the elements in order of increasing atomic mass

I put elements with similar properties into horizontal rows. I also left gaps incase any new elements were discovered

Dmitri Mendeleev



He

4 ← **MASS NUMBER** = number of protons + number of neutrons

2 ← **PROTON NUMBER** = number of protons (atomic)

CHARGE

N 14 / 7

Electron structure + 2,5

Mg 24 / 12

Electron structure + 2,8,2

Ca 40 / 20

Electron structure + 2,8,8,2

Isotopes

Arrange in a table with a different number of neutrons

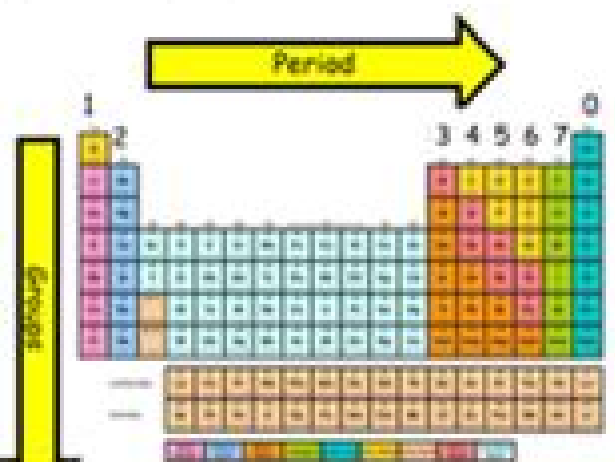
Notice that the mass number is different. How many neutrons does each isotope have?

16

17

18

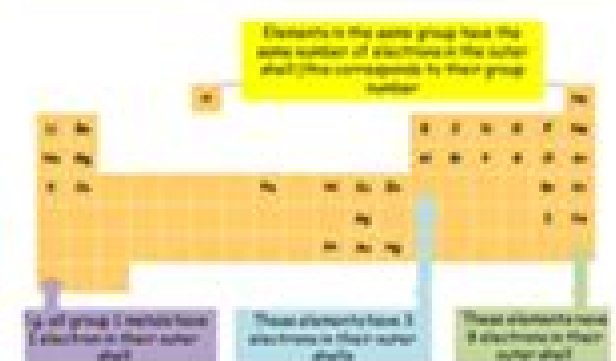
Each isotope has 8 protons - if it didn't then it just wouldn't be oxygen any more



How can an atom have a *decimal* for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 so the average atomic mass is:

$35 \times 3 + 1 \times 37 = 35.5$



(Most important) All of the elements in the same group have similar **PROPERTIES**. This is how I thought of the periodic table in the first place. This is called **PERIODICITY**.

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Emilie Sanchez

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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