



# **Explorer® 4250™**

## **Home Gateway**

### **User's Installation Guide**

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# Manual For Scientific Atlanta 4250

**Tao Wei**



## **Manual For Scientific Atlanta 4250:**

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1. The book is structured into several chapters, namely:
    - Chapter 1: Introduction to Manual For Scientific Atlanta 4250
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    - Chapter 4: Manual For Scientific Atlanta 4250 in Specific Contexts
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  6. In chapter 5, this book will draw a conclusion about Manual For Scientific Atlanta 4250. The final chapter will summarize the key points that have been discussed throughout the book.
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web understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from ph data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral ph at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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$\text{H}_2\text{SO}_4$  by titrating it with  $\text{NaOH}$  solution

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web solution a titrant volume 0 ml the solution pH is due to the acid ionization of  $\text{HCl}$  because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the pH of the solution is then  $\text{pH} = -\log 0.100 = 1.000$  b titrant volume 12.50 ml

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