

Test for Anions :

Anion	Test	Result
Carbonate (CO_3^{2-})	Add dilute acid (e.g. dilute hydrochloric acid)	Effervescence, CO_2 produced (turns limewater milky)
Sulfite (SO_3^{2-})	Add dilute acid (e.g. dilute hydrochloric acid), warm	SO_2 produced (turns aq. potassium manganate (VII) from purple to colourless)
Chloride (Cl^-) [in solution]	Acidify, then add aq. silver nitrate [Ag^+] Acidify, then add aq. lead nitrate [Pb^{2+}]	White ppt. of AgCl $\text{Ag}^+ (\text{aq}) + \text{Cl}^- (\text{aq}) \rightarrow \text{AgCl} (\text{s})$ White ppt. of PbCl_2
Bromide (Br^-) [in solution]	Acidify, then add aq. silver nitrate	Cream ppt. of AgBr
Iodide (I^-) [in solution]	Acidify, then add aq. silver nitrate (or lead nitrate)	Yellow ppt. of AgI (or PbI_2)
Sulfate (SO_4^{2-}) [in solution]	Acidify, then add aq. barium nitrate (or barium chloride) [Ba^{2+}]	White ppt. of BaSO_4 $\text{Ba}^{2+} + \text{SO}_4^{2-} \rightarrow \text{BaSO}_4$
Nitrate (NO_3^-)	Add aq. sodium hydroxide, then little aluminium (foil) and warm carefully	Ammonia produced (turns damp red litmus paper or universal indicator paper blue)

June 2002 Paper 6 Chemistry Igcse

Rachel Sandford



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Decoding **June 2002 Paper 6 Chemistry Igcse**: Revealing the Captivating Potential of Verbal Expression

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