

A chemical bond - forms when 2 or more atoms rearrange valence
electrons to increase stability.

ionic bond - forms when valence electrons are transferred
from one atom to another

cation - atom loses electrons to become + charged

anion - atom gains electrons to become - charged

In ionic compounds the ions are arranged in a crystal lattice.
Strong forces hold the ions together.

Properties of ionic compounds:

- high melting and boiling points
- hard - not easily crushed
- conduct electricity when melted or
dissolved because the ions are free to move.

covalent bond - electrons are shared, forming
molecules.

Covalent compounds have weaker forces holding the
molecules together.

Properties of covalent compounds:

- lower melting and boiling points
- Many covalent compounds are volatile liquids or gases.
- soft - easier to crush
- are not conductors of electricity

electronegativity - property that tells how strong an atom's attraction
is for electrons.

Since oxygen has a higher electronegativity than hydrogen, oxygen
holds onto shared electrons tighter, giving the oxygen a partial
negative charge and the hydrogen a partial positive charge.

Ionic Bonds And Notetaking Guide Answer

Yijin Wang



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Table of Contents Ionic Bonds And Notetaking Guide Answer

1. Understanding the eBook Ionic Bonds And Notetaking Guide Answer
 - The Rise of Digital Reading Ionic Bonds And Notetaking Guide Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds And Notetaking Guide Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonds And Notetaking Guide Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds And Notetaking Guide Answer
 - Personalized Recommendations
 - Ionic Bonds And Notetaking Guide Answer User Reviews and Ratings
 - Ionic Bonds And Notetaking Guide Answer and Bestseller Lists
5. Accessing Ionic Bonds And Notetaking Guide Answer Free and Paid eBooks
 - Ionic Bonds And Notetaking Guide Answer Public Domain eBooks
 - Ionic Bonds And Notetaking Guide Answer eBook Subscription Services
 - Ionic Bonds And Notetaking Guide Answer Budget-Friendly Options
6. Navigating Ionic Bonds And Notetaking Guide Answer eBook Formats

- ePub, PDF, MOBI, and More
- Ionic Bonds And Notetaking Guide Answer Compatibility with Devices
- Ionic Bonds And Notetaking Guide Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonds And Notetaking Guide Answer
 - Highlighting and Note-Taking Ionic Bonds And Notetaking Guide Answer
 - Interactive Elements Ionic Bonds And Notetaking Guide Answer
- 8. Staying Engaged with Ionic Bonds And Notetaking Guide Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonds And Notetaking Guide Answer
- 9. Balancing eBooks and Physical Books Ionic Bonds And Notetaking Guide Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonds And Notetaking Guide Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Bonds And Notetaking Guide Answer
 - Setting Reading Goals Ionic Bonds And Notetaking Guide Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Bonds And Notetaking Guide Answer
 - Fact-Checking eBook Content of Ionic Bonds And Notetaking Guide Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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