

### Percent Yield Problems

**Theoretical Yield** = The *maximum* amount of product that can be formed from given reactants.

--Theoretical yield is the answer from the stoichiometry calculation = maximum possible amount

---Number based upon the "theoretical" mathematical equation – no experiments were conducted

**Actual Yield** = The amount of product that *actually* forms when the reaction is carried out in the lab.

--Due to a variety of factors, the actual amount will ALWAYS be less than the theoretical yield.

---In real life, we never get as much as we anticipate, this "actual" amount from the lab experiment

**Percent Yield** = The *ratio* of the actual yield to the theoretical yield expressed as a percentage.

$$\text{Percent Yield} = \frac{\text{Actual}}{\text{Theoretical}} \times 100$$

**EXAMPLE)** How many grams of hydrofluoric acid (HF) are required to react completely with 23.68 grams of calcium hydroxide in the following reaction?  $\square$



$$23.68 \text{ g Ca(OH)}_2 \times \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} \times \frac{2 \text{ mol HF}}{1 \text{ mol Ca(OH)}_2} \times \frac{20.01 \text{ g HF}}{1 \text{ mol HF}} = \boxed{12.79 \text{ g HF}}$$

Based on the stoichiometry of this problem, the **Theoretical Yield** for this problem is 12.79 g HF. But, when this experiment was conducted, an **Actual Yield** of only 10.41 g HF was collected. Using the equation for percent yield, the **Percent Yield** of this experiment was 81.39%.

$$\text{Percent Yield} = \frac{\text{Actual}}{\text{Theoretical}} \times 100 = \frac{10.41 \text{ g HF}}{12.79 \text{ g HF}} \times 100 = \boxed{81.39\%}$$

1) A student adds 200.0g of  $\text{C}_7\text{H}_6\text{O}_3$  to an excess of  $\text{C}_2\text{H}_5\text{O}_2$ , this produces  $\text{C}_9\text{H}_8\text{O}_4$  and  $\text{C}_2\text{H}_4\text{O}_2$ . Calculate the percent yield if 231 g of aspirin ( $\text{C}_9\text{H}_8\text{O}_4$ ) is produced in an experiment.



2) According to the following equation, Calculate the percentage yield if 550.0 g of toluene ( $\text{C}_7\text{H}_8$ ) added to an excess of nitric acid ( $\text{HNO}_3$ ) provides 305 g of the p-nitrotoluene ( $\text{C}_7\text{H}_7\text{NO}_2$ ) product in a lab experiment.



3) Aluminum reacts with an aqueous solution containing excess copper (II) sulfate. If 1.85 g Al reacts and the percentage yield of Cu is 56.6%, what mass of Cu is produced?



4) The combustion of methane( $\text{CH}_4$ ) produces carbon dioxide and water. Assume that 2.0 mol of  $\text{CH}_4$  burned in the presence of excess air. What is the percentage yield if in an experiment the reaction produces 87.0 g of  $\text{CO}_2$ ?

5) 15.3 grams of Lithium is dropped into a solution containing excess copper II phosphate. When the reaction is completed, 1.25 grams of copper is formed. What is the percent yield?

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# Guided Study Work Answers Stoichiometry

**Michael Seilmaier**



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