

Section 7.2: Equilibrium Law and the Equilibrium Constant

Tutorial 1 Practice, page 431



Products: $2 \text{CO}(\text{g})$; $\text{O}_2(\text{g})$

Reactant: $2 \text{CO}_2(\text{g})$

$$\text{Equilibrium law equation: } K = \frac{[\text{CO}(\text{g})]^2 [\text{O}_2(\text{g})]}{[\text{CO}_2(\text{g})]^2}$$



Products: $4 \text{HCl}(\text{g})$; $\text{O}_2(\text{g})$

Reactants: $2 \text{Cl}_2(\text{g})$; $2 \text{H}_2\text{O}(\text{g})$

$$\text{Equilibrium law equation: } K = \frac{[\text{HCl}(\text{g})]^4 [\text{O}_2(\text{g})]}{[\text{Cl}_2(\text{g})]^2 [\text{H}_2\text{O}(\text{g})]^2}$$



Product: $3 \text{O}_2(\text{g})$

Reactant: $2 \text{O}_3(\text{g})$

$$\text{Equilibrium law equation: } K = \frac{[\text{O}_2(\text{g})]^3}{[\text{O}_3(\text{g})]^2}$$



Products: $2 \text{N}_2(\text{g})$; $6 \text{H}_2\text{O}(\text{g})$

Reactants: $4 \text{NH}_3(\text{g})$; $3 \text{O}_2(\text{g})$

$$\text{Equilibrium law equation: } K = \frac{[\text{N}_2(\text{g})]^2 [\text{H}_2\text{O}(\text{g})]^6}{[\text{NH}_3(\text{g})]^4 [\text{O}_2(\text{g})]^3}$$

2. **Given:** $\text{CO}(\text{g}) + 2 \text{H}_2(\text{g}) \rightleftharpoons \text{CH}_3\text{OH}(\text{g})$; $[\text{CO}(\text{g})]_{\text{equilibrium}} = 0.079 \text{ mol/L}$;

$[\text{H}_2(\text{g})]_{\text{equilibrium}} = 0.158 \text{ mol/L}$; $[\text{CH}_3\text{OH}(\text{g})]_{\text{equilibrium}} = 0.021 \text{ mol/L}$

Required: K

Solution: Write the equilibrium law equation using the balanced chemical equation. Then, substitute the equilibrium concentrations into the equilibrium law equation and solve for K .

$$\begin{aligned} K &= \frac{[\text{CH}_3\text{OH}(\text{g})]}{[\text{CO}(\text{g})][\text{H}_2(\text{g})]^2} \\ &= \frac{(0.021)}{(0.079)(0.158)^2} \end{aligned}$$

$$K = 11$$

Statement: The equilibrium constant, K , for the reaction is 11.

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Table of Contents Grade 12 Nelson Chemistry Questions Answers

1. Understanding the eBook Grade 12 Nelson Chemistry Questions Answers
 - The Rise of Digital Reading Grade 12 Nelson Chemistry Questions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 12 Nelson Chemistry Questions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 12 Nelson Chemistry Questions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 12 Nelson Chemistry Questions Answers
 - Personalized Recommendations
 - Grade 12 Nelson Chemistry Questions Answers User Reviews and Ratings
 - Grade 12 Nelson Chemistry Questions Answers and Bestseller Lists

5. Accessing Grade 12 Nelson Chemistry Questions Answers Free and Paid eBooks
 - Grade 12 Nelson Chemistry Questions Answers Public Domain eBooks
 - Grade 12 Nelson Chemistry Questions Answers eBook Subscription Services
 - Grade 12 Nelson Chemistry Questions Answers Budget-Friendly Options
6. Navigating Grade 12 Nelson Chemistry Questions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Grade 12 Nelson Chemistry Questions Answers Compatibility with Devices
 - Grade 12 Nelson Chemistry Questions Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 12 Nelson Chemistry Questions Answers
 - Highlighting and Note-Taking Grade 12 Nelson Chemistry Questions Answers
 - Interactive Elements Grade 12 Nelson Chemistry Questions Answers
8. Staying Engaged with Grade 12 Nelson Chemistry Questions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 12 Nelson Chemistry Questions Answers
9. Balancing eBooks and Physical Books Grade 12 Nelson Chemistry Questions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 12 Nelson Chemistry Questions Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Grade 12 Nelson Chemistry Questions Answers
 - Setting Reading Goals Grade 12 Nelson Chemistry Questions Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade 12 Nelson Chemistry Questions Answers
 - Fact-Checking eBook Content of Grade 12 Nelson Chemistry Questions Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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