

Types of Reactions

Goal

- group chemical changes into four basic types

Notes

★ Direct combination (synthesis)

Δ General Pattern



Δ Examples

- ⊛ Haber Process



- ⊛ process for making ammonia

- ⊛ developed in 1913 by a German scientist Fritz Haber

- ⊛ Ammonia used for making fertilizer and explosives

- ⊛ Combustion



★ Decomposition

Δ General Pattern



Δ Example



★ Single Replacement (substitution)

Δ General Pattern



Δ Examples

- ⊛ Replacement of a metal by a more active metal

Molecular equation



ionic equation



net ionic equation (striking spectator ions)



- ⊛ Replacement of a compound by a more active compound



★ Double Replacement (Exchange of Ions)

Δ General Pattern



Δ Example



★ Reversible and catal reactions

Δ Final reactions - reaction in which the products are not available to react to form the initial reactants because

- ⊛ A precipitate forms
- ⊛ A gas forms
- ⊛ A liquid forms
- ⊛ A product is removed

Δ Reversible reactions - reaction in which the products remain available to react to form the initial reactants because

- ⊛ They remain aqueous



Answer the questions below by circling the number of the correct response

Evans Chemistry Corner Final Paper An

Mike Jess



Evans Chemistry Corner Final Paper An:

The book delves into Evans Chemistry Corner Final Paper An. Evans Chemistry Corner Final Paper An is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Evans Chemistry Corner Final Paper An, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

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- Chapter 2: Essential Elements of Evans Chemistry Corner Final Paper An
- Chapter 3: Evans Chemistry Corner Final Paper An in Everyday Life
- Chapter 4: Evans Chemistry Corner Final Paper An in Specific Contexts
- Chapter 5: Conclusion

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